

Unit 2 Chemistry

Solubility Experiment - Deriving a Solubility Curve

Introduction

Solubility is defined as the measure of the amount of solute dissolved in a given amount of solvent at a given temperature.

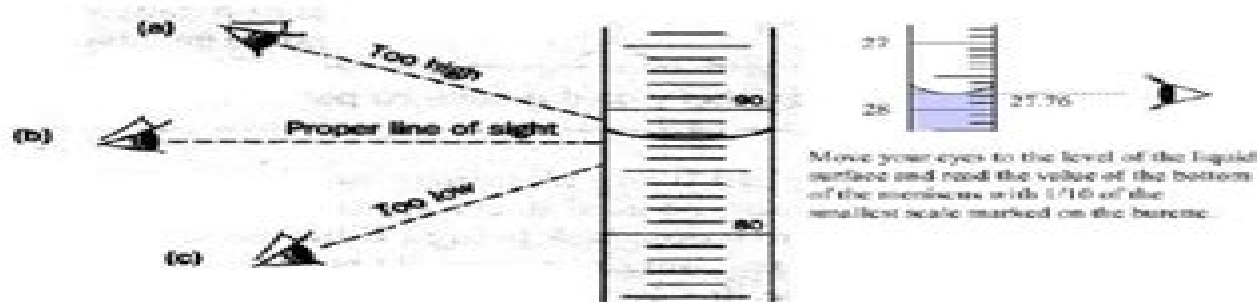
The solubility of ionic substances in water is generally measured as $\frac{\text{grams solute}}{100 \text{ g water}}$ at a given temperature.

A solubility curve is a graph of solubility (g/100g water) versus temperature for a particular solute in a particular solvent.

In this practical exercise we will be determining the solubility of potassium nitrate in g/100g water at various temperatures and graphing the results to derive a solubility curve.

In our calculations we will assume that the density of water is 1.0 g/mL at all temperatures. When we add 5.0mL of water we assume that this is equivalent to adding 5.0g of water.

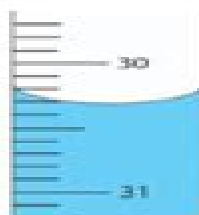
Accurate measurement of mass and volume is important in this practical exercise. Record all four decimal places from the electronic balance and read the measurements for volume from a horizontal line of sight to the burette, using the bottom of the meniscus.



Reference: Ch 13.3 & Ch 14.1 Heinemann Chemistry 1

Pre-lab activity

1. Record the volume measurement for this diagram.



Deriving A Solubility Curve Prac

A. S. Paintal, Pritam Gill-Kumar

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the authors' names.

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