

Name _____

Section _____

Date _____

Post-Laboratory Questions

(Use the spaces provided for the answers and additional paper if necessary.)

1. A student doing the experiment in this module extended the study of the $[\text{CoCl}_4]^{2-}$ ion/ $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ ion equilibrium. In one test tube, the student added silver nitrate (AgNO_3) to a blue equilibrium mixture prepared from CoCl_2 . The reaction mixture became pink and cloudy, and a white precipitate settled out leaving a clear, pink solution. The student identified the precipitate as silver chloride, AgCl .

(1) Briefly explain how these observations are consistent with Le Châtelier's principle.

In a second test, the student placed a test tube containing a pink equilibrium mixture in a hot-water bath. The solution turned blue. When the student removed the test tube from the hot-water bath and placed it in an ice-water bath, the solution turned pink.

(2) Is the forward reaction in the $[\text{CoCl}_4]^{2-}$ ion/ $[\text{Co}(\text{H}_2\text{O})_6]^{2+}$ ion equilibrium exothermic or endothermic?

(3) Briefly explain how the student's observations support your answer to (2).

(4) Write the net ionic equation for this equilibrium, including heat.

Data Sheet Equil0392 Introducing Equilibrium With Answers

AW Rasmussen



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Metals Abstracts ,1993

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