

Name key

Physical Science Chemical Bonding Study Guide

A. Covalent & Ionic Compounds/Bonds

1. When an atom loses an electron it becomes a positive ion (cation), and when it gains an electron it becomes a negative ion (anion).
2. An element in the halogen group has 7 valence electrons, and wants to gain 1 more to have 8. This is because of the octet rule.
3. An ionic bond is the attraction between oppositely charged ions, AKA a metal and a nonmetal.
4. In a covalent bond, atoms share electrons, whereas in an ionic bond, atoms transfer electrons.
5. (Honors only) When electrons are shared unequally in a covalent bond it is polar, and when they're shared equally it is nonpolar. You can predict if a compound is ionic, polar covalent or nonpolar covalent by subtracting the electronegativities.
6. An atom can have up to 8 valence electrons. The family (column) on the periodic table that all have 8 valence electrons is called noble gases.
7. Atoms can gain, lose, or share valence electrons to become more stable.

B. Formulas for Ionic compounds - Fill in the chart below

	Magnesium	Fluorine	Nitrogen	Lithium
# OF VALENCE ELECTRONS	<u>2</u>	<u>7</u>	<u>5</u>	<u>1</u>
OXIDATION #	<u>+2</u>	<u>-1</u>	<u>-3</u>	<u>+1</u>

2. What is the **chemical formula** for each of the following ionic compounds?
(*hint: you can either draw the Lewis structures and use arrows to show the transfer of electrons, OR use the periodic table to find their oxidation numbers. They will combine so that their oxidation numbers cancel out to zero. Use the criss cross method if their oxidation numbers do not cancel.*)

- a) Potassium and Chlorine - KCl
- b) Calcium and Bromine - CaBr₂
- c) Magnesium and oxygen - MgO

Chemical Bonds Study Guide Answer Key

Barbara Herlihy



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Table of Contents Chemical Bonds Study Guide Answer Key

1. Understanding the eBook Chemical Bonds Study Guide Answer Key
 - The Rise of Digital Reading Chemical Bonds Study Guide Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemical Bonds Study Guide Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an eBook Platform
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemical Bonds Study Guide Answer Key
 - Personalized Recommendations
 - Chemical Bonds Study Guide Answer Key User Reviews and Ratings
 - Chemical Bonds Study Guide Answer Key and Bestseller Lists

5. Accessing Chemical Bonds Study Guide Answer Key Free and Paid eBooks
 - Chemical Bonds Study Guide Answer Key Public Domain eBooks
 - Chemical Bonds Study Guide Answer Key eBook Subscription Services
 - Chemical Bonds Study Guide Answer Key Budget-Friendly Options
6. Navigating Chemical Bonds Study Guide Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemical Bonds Study Guide Answer Key Compatibility with Devices
 - Chemical Bonds Study Guide Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemical Bonds Study Guide Answer Key
 - Highlighting and Note-Taking Chemical Bonds Study Guide Answer Key
 - Interactive Elements Chemical Bonds Study Guide Answer Key
8. Staying Engaged with Chemical Bonds Study Guide Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemical Bonds Study Guide Answer Key
9. Balancing eBooks and Physical Books Chemical Bonds Study Guide Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemical Bonds Study Guide Answer Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemical Bonds Study Guide Answer Key
 - Setting Reading Goals Chemical Bonds Study Guide Answer Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemical Bonds Study Guide Answer Key
 - Fact-Checking eBook Content of Chemical Bonds Study Guide Answer Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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