

My "Mole Concept Test" will take place on: _____!

Before I write my "Mole Concept Test", I will be able to:

- ☐ **explain why the mole is used in chemistry**
 - know how to find *atomic mass* on the periodic table
 - know the relationship between atomic mass and molar mass
 - recognize that a mole is a unit for counting atoms, ions, and molecules
- ☐ **perform calculations involving the mole**
 - determine the molar mass of an element or compound
 - convert among particles, moles, and mass
- ☐ **understand the relationship between moles and gases at STP**
 - know the volume of a mole of gas at STP (i. e. 22.4 L)
 - convert among moles, mass, and volume of a gas at STP
- ☐ **perform calculations involving molecular and empirical formulae**
 - distinguish between molecular and empirical formulae
 - determine the percent composition by mass from the formula of a compound
 - determine the empirical formula for a compound from its percent composition by mass
 - determine a molecular formula from the molecular mass and empirical formula
- ☐ **Describe concentration in terms of molarity**
 - Know that the units for molarity/concentration are mol/L or M
 - Write a procedure for how to prepare a solution of known molarity (a standard solution)
- ☐ **Perform calculations involving molarity**
 - Perform calculations relating mass, moles, volume, and molarity
- ☐ **Understand the meanings of the following vocabulary words:**
 - atomic mass
 - empirical formula
 - molarity
 - molar mass
 - molar solution
 - mole
 - molecular formula
 - molecular mass
 - quantitative transfer
 - percentage composition
 - standard solution
 - STP

Chemistry Holt Study Guide Mole Concept

Thomas Griffiths



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