

APHY101- Anatomy & Physiology I

Chapter #2- Chemistry Comes Alive

Study Guide

- Define matter. What is the smallest unit of matter?

Matter = Anything that occupy space and has mass. Can exist in different forms: Solid, Liquid, or Gas.

- What is an element?

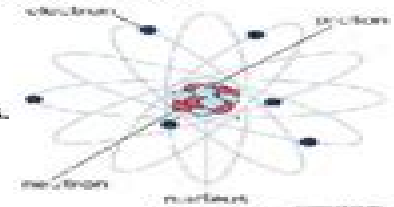
Element = a pure substance that cannot be created or broken down into simpler substances by ordinary chemical means. There're 118 recognized elements and 92 occur in nature, ex. Carbon, Oxygen, Hydrogen, Nitrogen (found most in human body). Every element is made up of building blocks called **ATOMS**.

- Define atomic number, atomic mass & atomic weight.

Atomic Number = the number of protons in the nucleus of an atom.

Atomic Mass = the sum of the number of neutrons and protons in an atomic nucleus.

Atomic Weight = The average mass of all the isotopes of an element.

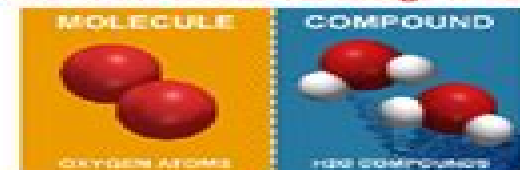


- Define atom

Atom = Smallest units of matter, no life properties (non-living), made up for every tiny Subatomic Particles.

- What are the three (3) subatomic particles that make up an atom & what is the location & charge of each?

- Nucleus** – Central core of an atom.
 - Protons (+): charged particles
 - Neutrons: Neutral; not charged
- Electrons** (-): charged particles that orbit the nucleus



- Define molecule

Molecule = A group of Atoms bonded together by chemical bonds. Can contain only one element or more.

- Define compound

Compound = Substance composed of two or more elements joined by chemical bonds.

For example, water (H₂O) has three atoms, two hydrogen (H) atoms and one oxygen (O) atom.

- All compounds are molecules, but not all molecules are compounds*



- Know the difference between an organic molecule/compound & an inorganic molecule/compound

Organic compounds always contain carbon while most inorganic compounds do not contain carbon.

- Why is water an important inorganic compound in the body?

Water is the most abundant inorganic compound, making up over 60% of the volume of cells and over 90% of body fluids like blood. Many substances dissolve in water and all the chemical reactions that take place in the body do so when dissolved in water. Other inorganic molecules help keep the acid/base balance (pH) and concentration of the blood and other body fluids stable.

Chemistry Comes Alive Study Guide

**Yash Gupta, Vishal Rai, Soban
Khan, Nisha Bano**



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Student Guide For Living Chemistry Barbara Ucko, 2012-12-02 Student Guide for Living Chemistry is a 23 chapter textbook guide that allows students to study and review on their own and test their understanding to help them prepare for examinations Every chapter begins with a list of objectives stating exactly the skills to develop in a particular unit Each objective corresponds to a section in the textbook Living Chemistry Three kinds of questions are provided for each objective to check the student's understanding namely short answer Study Questions multiple choice and fill in The answers for all questions are provided at the end of the chapter The opening chapters cover the SI units composition of matter chemical bonding compounds chemical change gases respiration and water The subsequent chapters deal with solutions acids bases salts nuclear and organic chemistry oxygen derivatives and hydrocarbons polymers and other organic derivatives This textbook also explores the chemistry of carbohydrates lipids proteins enzymes and energy and carbohydrate metabolism The remaining chapters discuss the chemistry of vitamins hormones body fluid drugs and poisons Undergraduate chemistry students will find this book invaluable

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Laboratory Units in Chemistry William McPherson,William Edwards Henderson,George Winegar Fowler,1938

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