Chemistry 105 Equation Sheet

Constants

 $a_0 = 0.5291772059 \text{ Å}$ $c = 2.99792458 \times 10^{\circ} \,\mathrm{m/s}$ $e = 1.6021765 \times 10^{-15} C$ $\epsilon_0 = 8.8541877 \times 10^{-12} \text{ F/m}$ $h = 6.626 \times 10^{-14} \text{ J·s}$ $k_0 = 1.38065 \times 10^{-23} \text{ J/K}$ $k_s = 8.987551 \times 10^9 \frac{N \cdot m^2}{c^2}$ $m_e = 9.109383 \times 10^{-31} \text{ kg}$ $m_n = 1.6749273 \times 10^{-27} \text{ kg}$ $m_p = 1.6726217 \times 10^{-27} \text{ kg}$ $N_A = 6.022142 \times 10^{21} \text{ mo}^{12}$ $R = 8.31447 \frac{J}{mod - K}$ = 0.0820575 \(\frac{L}{mol \cdot K}\)
= 1.987207 \(\frac{cal}{mol \cdot K}\) $R_{\rm H} = 1.0974 \times 10^7 \, {\rm m}^{-1}$

Conversions (* = exact)

Length

 $1 \text{ in} = 2.54 \text{ cm}^*$ $1 \text{ ft} = 12 \text{ in ches}^*$ $1 \text{ vd} = 3 \text{ ft}^*$ $1 \text{ m} = 39.37 \text{ in ches}^*$ 1mile = 5280 feet* 1mile = 1.609 kmMass 1 lb = 453.59237 g16 oz. avdp.*

1 US lb = 12 troy ounces * 1 ton = 2000 lb * 1 amu = 1.660533 × 10⁻²⁴ g

Volume

 $1 \text{ in}^3 = 16.4 \text{ cm}^3$

1 US gal = 3.7854 L = 4 US qt * 1 US at = 32 US fl. ozs. * 1 US fluid oz. = 29.5735 mL. $1 \text{ cm}^3 = 0.001 \text{ L}^*$

Pressure

1 atm = 760 torr * = 101325 Pa * = 406.783 in H₂O $= 14.6959 \text{ lb/in}^2 \text{ (psi)}$ 1 torr = 1 mm Hg (at 0 °C)* 1 Pa = 1 kg mr 1 s 2* 1 bar = 100,000 Pa *

Force:

 $1 \text{ N} = 1 \text{ kg m s}^{2}$

Power.

1 Watt = 1 L/s

Emergy

 $1 \text{ J} = 1 \text{ kg/m}^2/\text{s}^2 = 0.2390 \text{ cal}$ 1 J = 1 V·C = 0.009869 L·atm $1 \text{ eV} = 1.602 \times 10^{-19} \text{ I}$ 1 L-atm = 101.3 J

Formulas.

 $\rho = m/V$ $T_F = \frac{9^{np}}{500} T_C + 32 \, ^{np}$ $T_K = \frac{1K}{100}T_C + 273.15 \text{ K}$ $c = \lambda \nu$ $E = h\nu$ $\lambda = \frac{h}{mv}$ Example $= V_2 \text{ min}^2$ $E = mc^c$ $\Phi = h\nu \cdot KE_{\text{theorem}}$

 $E_m = -Z^2 \frac{R_H hc}{r}$

 $\Delta x \Delta(mv) \ge \frac{h}{4\pi}$ $\frac{1}{2} = R_H \left(\frac{1}{n_c^2} - \frac{1}{n_c^2} \right)$ $V(\mathbf{r}) = \frac{Q_1 Q_2}{4\pi r_1 r_2}$ $F(r) = k_s \frac{Q_1 Q_2}{r}$ u = 0r $\Delta E = q + w$ $\Delta E_{max} = 0$ $\Delta S_{univ} \ge 0$ $S = k_B \ln W$ $q = mC\Delta T$ $w = -P_{ex} \Delta V$ $\Delta H = \Delta E + P\Delta V$ $\Delta S = \frac{q_{rev}}{\pi}$ $\Delta G = \Delta H - T\Delta S$ $\Delta G = \Delta G^{\circ} + RT \ln O$ $\Delta H_{ryn} = \Sigma n H_{f,nead} - \Sigma m H_{f,nead}$ $\Delta H'_{max} \approx \Sigma n D_{maximum}$ $-\Sigma mD_{formed}$ PV = nRT $P_{tot} = P_1 + P_2 + P_3 + \dots$ $\frac{Rate_A}{Rate_B} = \frac{Time_B}{Time_A} = \sqrt{\frac{M_B}{M_A}}$

$$\frac{Rate_A}{Rate_B} = \frac{Time_B}{Time_A} = \sqrt{\frac{M_B}{M_A}}$$

$$KE_{avs}=\frac{3}{2}RT$$

$$u_{rms} = \sqrt{v^2} = \sqrt{\frac{3RT}{M}}$$

$$pX = -\log X$$

$$K_{er} = [H_3O^+][OH]$$

^{* =} exact v = frequency v = velocity

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Chemistry Equation Sheet

Princeton Review

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