

Study Guide for Chapter 7 – Chemical Quantities

Test Setup: Multiple Choice, Matching, Problems: Find molar mass, determine % composition, Molar conversions will include all of the following: moles-particles, moles-mass, and moles-volume, determine the empirical formulas & molecular formulas by looking at actual formulas.

Sections in the Book:

7.1 – The Mole: A Measurement of Matter

Objectives:

- Describe how Avogadro's number is related to a mole of any substance
- Calculate the mass of a mole of any substance

Vocabulary

- mole (mol), Avogadro's number, representative particle

Practice Problems:

1. Find the molar mass of Li_2S .
2. Find the molar mass of $\text{Ca}(\text{OH})_2$.
3. How many moles in 1.50×10^{23} molecules NH_3 .
4. How many formula units in 6.78 moles of sodium chloride (NaCl)

7.2 – Mole-Mass and Mole-Volume Relationships

Objectives:

- Use the molar mass to convert between mass and moles of a substance.
- Use the mole to convert among measurements of mass and volume of substances.

Vocabulary:

- molar mass, standard temperature and pressure (STP), molar volume

Practice Problems:

1. What is the mass in grams of 0.160 mol H_2O_2 ?
2. How many moles are in 508.4 g $\text{Ca}(\text{NO}_3)_2$?
3. Calculate the volume, in liters, of 8.75 mol of chlorine gas at STP.
4. How many moles are in a container filled with 45.9 L of oxygen gas at STP?

7.3 – Percent Composition and Chemical Formulas

Objectives:

- Calculate the percent composition of a substance from its chemical formula or experimental data.
- Derive the empirical formula and the molecular formula of a compound from experimental data.

Chemistry Chemical Quantities Study Guide

Michael Seilmaier



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