

Acid-Base Titrations Pre-Lab Questions (10 Points)

Name _____ Lab Day _____

Lab Instructor _____ Date _____

Show calculations in space provided or on a separate sheet of paper.

1. In an experiment, $\overset{\sqrt{1}}{39.26 \text{ mL}}$ of $\overset{\sqrt{1}}{0.1062 \text{ M}}$ NaOH solution was required to titrate $\overset{\sqrt{2}}{37.54 \text{ mL}}$ of unknown acetic acid solution to a phenolphthalein end point. Calculate the molarity of the acetic acid solution, and the percent (by weight) of acetic acid in the solution (assuming its density to be 1.00 g/mL).

$$\begin{aligned} (M_2)(V_2) &= (M_1)(V_1) \\ M_2(37.54 \text{ mL}) &= (0.1062 \text{ M})(39.26 \text{ mL}) \\ M_2 &= \frac{4.1654 \text{ M}}{37.54 \text{ mL}} = 0.1110 \dots \end{aligned}$$

0.1111 M 11.11 %

2. What is the proper disposal of all chemicals in this experiment?

Discard all solutions, HCl solution, rinsing solutions, and titration solutions into your waste beaker. If your waste beaker gets filled, discard all solutions in the waste-acid base container.

3. List some the hazards associated with sodium hydroxide.

Sodium hydroxide is toxic ~~and~~, corrosive, and can easily absorbed by the skin.

Chemfax Acid Base Titrations Lab Prelab Answers

Robert De Levie



Chemfax Acid Base Titrations Lab Prelab Answers:

Acid-base Titrations in Nonaqueous Solvents James Sherwood Fritz, 1952 **Identification of Misconceptions Concerning Acid-base Titrations** Lockwood Ryerson Gibbs, 1995 **Aqueous Acid-base Equilibria and Titrations** Robert De Levie, 1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an out moded way This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved **Acid/Base Titrations** Michael O'Brien, 1999 **Numerical Evaluation of Potentiometric Acid-base Titrations** Sten Johansson, 1978 **pH, Buffers and Acid-base Titrations** LaMotte Chemical Products Co. Educational Products Division, 1977

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1. Understanding the eBook Chemfax Acid Base Titrations Lab Prelab Answers
 - The Rise of Digital Reading Chemfax Acid Base Titrations Lab Prelab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemfax Acid Base Titrations Lab Prelab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemfax Acid Base Titrations Lab Prelab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemfax Acid Base Titrations Lab Prelab Answers
 - Personalized Recommendations
 - Chemfax Acid Base Titrations Lab Prelab Answers User Reviews and Ratings
 - Chemfax Acid Base Titrations Lab Prelab Answers and Bestseller Lists
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 - Chemfax Acid Base Titrations Lab Prelab Answers eBook Subscription Services
 - Chemfax Acid Base Titrations Lab Prelab Answers Budget-Friendly Options
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 - Chemfax Acid Base Titrations Lab Prelab Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Chemfax Acid Base Titrations Lab Prelab Answers
 - Interactive Elements Chemfax Acid Base Titrations Lab Prelab Answers
8. Staying Engaged with Chemfax Acid Base Titrations Lab Prelab Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemfax Acid Base Titrations Lab Prelab Answers
9. Balancing eBooks and Physical Books Chemfax Acid Base Titrations Lab Prelab Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemfax Acid Base Titrations Lab Prelab Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemfax Acid Base Titrations Lab Prelab Answers
- Setting Reading Goals Chemfax Acid Base Titrations Lab Prelab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemfax Acid Base Titrations Lab Prelab Answers
- Fact-Checking eBook Content of Chemfax Acid Base Titrations Lab Prelab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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