

Exercises

1. Use the vector notation (see Figure 2 above) to designate the dipole moments of:
CO ; HI ; ClF .
2. Write the Lewis structure for NO_3^- . Explain why the dipole moment is zero. (Hint: where does the center of negative charge have to be such that the dipole moment is zero?)
3. Classify each of the following molecules (or ions) as nonpolar (dipole moment = 0) or polar (dipole moment $\neq 0$). (Hint: it is generally a good idea to determine the shape of the molecule.)

O_2 ; I_2 ; CO ; NO ; H_2O ; CO_3^{2-} ; SO_4^{2-} ; SO_2



4. Which molecule of each pair has a dipole moment of zero?

a) N_2 or NO	b) CH_4 or CH_3Cl	c) SO_2 or SO_3
d) NO_3^- or NO_2^-	e) SO_4^{2-} or SO_2	f) CO_2 or OCS
5. Indicate whether the following statement is true or false and explain your reasoning.

The dipole moment of NH_3 is smaller than the dipole moment of CCl_4 .

6. Which molecule of each pair has the greater dipole moment?

- a) CCl_4 or CH_3Cl
- b) CH_3Br or CH_3Cl
- c) NF_3 or NI_3
- d) H_2O or H_2S

7. Based on the data in Table 1, predict the dipole moment of CH_3F .

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Xiang Xie



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