

PROCEDURE**CHECK ALL GLASSWARE!!!**

For this experiment, use the table below (Table 1) to help you prepare your mixtures.

Table 1. Reagents for determining the rate orders of MnO_4^- , $\text{H}_2\text{C}_2\text{O}_4$, and H^+ .

Reagents	Mixture 1	Mixture 2	Mixture 3	Mixture 4
0.50 M $\text{H}_2\text{C}_2\text{O}_4$	1.00 mL	1.00 mL	2.00 mL	1.00 mL
0.10 M H_2SO_4	1.00 mL	1.00 mL	1.00 mL	2.00 mL
Distilled H_2O	2.00 mL	3.00 mL	1.00 mL	1.00 mL
0.0050 M KMnO_4	2.00 mL	1.00 mL	2.00 mL	2.00 mL
Total	6.00 mL	6.00 mL	6.00 mL	6.00 mL

NOTE: Contamination can be a real problem for this experiment. Pipettes meant for a particular solution should be used only for that solution. The test tube should be clean and dry before kinetics run.

BEFORE STARTING ON PART A: Start your hot water bath for part B. Fill the beaker full of water.

Part A: Determining the Rate Law

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Chemistry 1212, LAB Manual Jack Duff, Alan Gabrielli, 2014-08-01 **Principles of Chemistry II** Georgia College & State University, 2025-08-15 Lab Manual for Chemistry 1212K Principles of Chemistry II **CHEM 1412 General Chemistry II Lab Manual**, 2022 **Acp Chem 1172 General Chemistry Lab Manual** Brooks/Cole, 2015-11-11
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