### **Isotope Practice**

1. Here are three isotopes of an element: 12C

VC. M/C

- a. The element is: Carbon
- b. The number 6 refers to the: Atomic Number
- The numbers 12, 13, and 14 refer to the different mass numbers
- d. How many protons and neutrons are in the first isotope? P: 6 N: 6
- e. How many protons and neutrons are in the second isotope? P: 6
- f. How many protons and neutrons are in the third isotope? P: 6 N: 8
- 2. Complete the following chart-

Isotope name	atomic#	mass #	# of protons	# of neutrons	# of electrons
Potassium-37	19	37	19	18	19
Oxygen-17	8	17	8	9	8
uranium-235	92	235	92	143	92
uranium-238	92	238	92	146	92
boron-10	5	10	5	5	5
boron-11	5	11	5	6	5

3. How do you calculate the mass number of an element?

Mass # = protons + neutrons

- 4. If an atom has 14 protons, 17 neutrons and 14 electrons, which expression identifies the element?
  - Silicon-14
  - b. Silicon-28
  - c. Silicon-31
  - d. Sillicon-17.
- 5. If an atom has 53 protons, 54 neutrons and 53 electrons, which expression identifies the element?
  - a. lodine-53
  - b. Iodine-127
  - c. Iodine-54
  - d. lodine-107

# **Compass Learning Oddyssey Answer Key Chemistry**

M Mark

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