



COVALENT BONDING

Practice Problems

Do your answers to these very practice problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a. Na_2O ionic b. CH_4 covalent c. HCl covalent
d. KNO_3 ionic e. H_2O covalent

2. Write the formula for the following molecular compounds.

- a. carbon dioxide b. sulfur dioxide c. carbon monoxide
d. nitrogen e. sulfur f. phosphorus

3. Write the Lewis structure for each of the following molecular compounds.

a. CH_4 b. H_2O c. CO_2 d. H_2S e. NH_3 f. CH_3Cl g. CH_3OH h. CH_3NH_2 i. $\text{CH}_3\text{CH}_2\text{CH}_3$ j. $\text{CH}_3\text{CH}_2\text{CH}_2\text{CH}_3$

4. Draw the Lewis structure for each of the following molecular compounds. Be sure to show all lone pairs of electrons and indicate the formal charge on each atom.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF .



2. Draw the Lewis structure for methane, CH_4 .



3. Draw the Lewis structure for carbon dioxide, CO_2 .



4. Draw the Lewis structure for acetylene, C_2H_2 .

5. Draw the Lewis structure for each of the following molecular compounds. Be sure to show all lone pairs of electrons and indicate the formal charge on each atom.



6. Draw the Lewis structure for each of the following molecular compounds. Be sure to show all lone pairs of electrons and indicate the formal charge on each atom.

SECTION 8.3 BONDING: THE ORBITAL

1. Draw the Lewis structure for each of the following molecular compounds. Be sure to show all lone pairs of electrons and indicate the formal charge on each atom.

2. Draw the Lewis structure for each of the following molecular compounds. Be sure to show all lone pairs of electrons and indicate the formal charge on each atom.

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Table of Contents Covalent Bonding Study Guide Answers Pearson Education

1. Understanding the eBook Covalent Bonding Study Guide Answers Pearson Education
 - The Rise of Digital Reading Covalent Bonding Study Guide Answers Pearson Education
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Study Guide Answers Pearson Education
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonding Study Guide Answers Pearson Education
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Study Guide Answers Pearson Education
 - Personalized Recommendations
 - Covalent Bonding Study Guide Answers Pearson Education User Reviews and Ratings
 - Covalent Bonding Study Guide Answers Pearson Education and Bestseller Lists
5. Accessing Covalent Bonding Study Guide Answers Pearson Education Free and Paid eBooks
 - Covalent Bonding Study Guide Answers Pearson Education Public Domain eBooks
 - Covalent Bonding Study Guide Answers Pearson Education eBook Subscription Services
 - Covalent Bonding Study Guide Answers Pearson Education Budget-Friendly Options
6. Navigating Covalent Bonding Study Guide Answers Pearson Education eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonding Study Guide Answers Pearson Education Compatibility with Devices
 - Covalent Bonding Study Guide Answers Pearson Education Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonding Study Guide Answers Pearson Education
 - Highlighting and Note-Taking Covalent Bonding Study Guide Answers Pearson Education
 - Interactive Elements Covalent Bonding Study Guide Answers Pearson Education
8. Staying Engaged with Covalent Bonding Study Guide Answers Pearson Education

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Covalent Bonding Study Guide Answers Pearson Education
- 9. Balancing eBooks and Physical Books Covalent Bonding Study Guide Answers Pearson Education
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding Study Guide Answers Pearson Education
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Covalent Bonding Study Guide Answers Pearson Education
 - Setting Reading Goals Covalent Bonding Study Guide Answers Pearson Education
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Covalent Bonding Study Guide Answers Pearson Education
 - Fact-Checking eBook Content of Covalent Bonding Study Guide Answers Pearson Education
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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