

Concentration and Molarity (M) table

Name: _____

Mass concentration (g/L)

Using the concentration and the formula for Molarity or the mass, calculate the molarity

Mass of substance	Volume (L)	Molarity of substance (M)	Mass of substance	Volume (L)	Molarity of substance (M)
100	10	10.00	10.000	10	1.00
20	20	1.00	20	10.0	2.00
1.0	10	0.10	1.0	0.1	1.00
20	20	1.00	54.4	0.1	54.4

Concentration (Moles/L) and Molarity (Moles/L) and Molarity (Moles/L) table

1. **Calculate the molarity of a solution made by dissolving 100g of NaCl in 10L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{100}{10} = 10.00 \text{ M}$$
2. **Calculate the molarity of a solution made by dissolving 20g of NaCl in 20L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{20}{20} = 1.00 \text{ M}$$
3. **Calculate the molarity of a solution made by dissolving 1.0g of NaCl in 10L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{1.0}{10} = 0.10 \text{ M}$$
4. **Calculate the molarity of a solution made by dissolving 20g of NaCl in 0.1L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{20}{0.1} = 200.0 \text{ M}$$
5. **Calculate the molarity of a solution made by dissolving 54.4g of NaCl in 0.1L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{54.4}{0.1} = 544.0 \text{ M}$$
6. **Calculate the molarity of a solution made by dissolving 100g of NaCl in 10L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{100}{10} = 10.00 \text{ M}$$
7. **Calculate the molarity of a solution made by dissolving 20g of NaCl in 20L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{20}{20} = 1.00 \text{ M}$$
8. **Calculate the molarity of a solution made by dissolving 1.0g of NaCl in 10L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{1.0}{10} = 0.10 \text{ M}$$
9. **Calculate the molarity of a solution made by dissolving 20g of NaCl in 0.1L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{20}{0.1} = 200.0 \text{ M}$$
10. **Calculate the molarity of a solution made by dissolving 54.4g of NaCl in 0.1L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{54.4}{0.1} = 544.0 \text{ M}$$
11. **Calculate the molarity of a solution made by dissolving 100g of NaCl in 10L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{100}{10} = 10.00 \text{ M}$$
12. **Calculate the molarity of a solution made by dissolving 20g of NaCl in 20L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{20}{20} = 1.00 \text{ M}$$
13. **Calculate the molarity of a solution made by dissolving 1.0g of NaCl in 10L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{1.0}{10} = 0.10 \text{ M}$$
14. **Calculate the molarity of a solution made by dissolving 20g of NaCl in 0.1L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{20}{0.1} = 200.0 \text{ M}$$
15. **Calculate the molarity of a solution made by dissolving 54.4g of NaCl in 0.1L of water.**

$$\text{Molarity} = \frac{\text{Mass}}{\text{Volume}} = \frac{54.4}{0.1} = 544.0 \text{ M}$$

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Answer Key

Huangqi Zhang



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