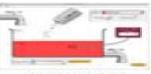
Concentration and Molarity II: Dilution and Evoporation PhET-Chemistry Labs

Introduction: In the first part of this lab, you learned about the actions of solid salts added to water to form solutions of various concentrations. In this exercise, you will use a powerful and simple formula to determine the concentration of a solution when it has been diluted with additional water or made more concentrated due to evaporation of water.



Concentration Some handy vocabulary for you to define: Molarity Dilution Evaporation Procedure: PhET→Play with the Sims → Chemistry → Concentration Past North $M_1V_1=M_2V_2$ Port 1: Dilution At this point you should be familiar with the Concentration simulation. Because some time may have passed, take a few minutes to re-learn the simulations. Pay particular attention to the effect of evaporation and addition of water to the solution's concentration How does the concentration change as additional water is added? * How does the concentration change as evaporation occurs? Why is this? Does evaporation change the concentration of a saturated solution? Why is this? Using the concentrated solution spigot, add a % Liter of Drink Mix to an empty beaker. What is the concentration? Is this solution saturated? How do you know? Fill the beaker with another % Liter of water. What is the new concentration? Complete the table below, using Cobalt chloride in an empty beaker, writing the concentration in the boxes provided. Only .25L of spigot solution .25L spigot + .25L water .25L spigot + .50 L water .25L spigot + .75 L water Repeat the exercise, using Potassium dichromate in an empty beaker. Only :25L of spigot solution .25L spigot+.25L water .25L spigot + .50 L water .25L spigot + .75 L water

What do you notice about the concentration change as each addition of .25L of water is added to the concentrated spigot solution?

Concentration And Molarity Ii Dilution And Evaporation

Stephen D. Cramer

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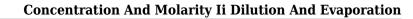
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