Stoichiometry

Section 11.1 Defining Stoichiometry

pages 368-372

Practice Problems

pages 371-372

- Interpret the following balanced chemical equations in terms of particles, moles, and mass.
 Show that the law of conservation of mass is observed.
 - a. $N_s(g) + 3H_s(g) \rightarrow 2NH_s(g)$

1 molecule N₂ + 3 molecules H₂ → 2 molecules NH₃

1 mole N₂ + 3 moles H₂ -> 2 moles NH₃

Milance

N₂: (2 mol)(14.007 g/mol) = 28.014 g

3H₂: (6 mol)(1.008 g/mol) = 6.048 g

2NH₃: (2 mol)(14.007 g/mol) + (6 mol)(1.008 g/mol) = 34.062 g

28.014 g N₂ + 6.048 g H₂ -> 34.062 g NH₃

34.062 g reactants = 34.062 g products

b. HCl(aq) + KOH(aq) → KCl(aq) + H₂O(I)

1 molecule HCl + 1 formula unit KOH → 1 formula unit KCl + 1 molecule H₂O

1 mole HCI + 1 mole KOH →

1 mole KCl + 1 mole H₂O

Politica execu-

HCl: (1 mol)(1.008 g/mol) + (1 mol)(35.453 g/mol) = 36.461 g

KOH: (1 mol)(39.098 g/mol) + (1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol) = 56.105 g

KCl: (1 mol)(39.098 g/mol) + (1 mol)(35.453 g/mol) = 74.551 g

 H_2O : (2 mol)(1.008 g/mol) + (1 mol)(15.999 g/mol) = 18.015 g

36.461 g HCl + 56.105 g KOH --74.551 g KCl + 18.015 g H₂O

92.566 g reactants = 92.566 g products

c. $2Mg(s) + O_s(g) \rightarrow 2MgO(s)$

2 atoms Mg + 1 molecule O₁ → 2 formula units MgO

2 moles Mg + 1 mole O₂ → 2 moles MgO

Marks

2Mg: (2 mol)(24.305 g/mol) = 48.610 g

O₂: (2 mol)(15.999 g/mol) = 31.998 g

2MgO: (2 mol)(24.305 g/mol) + (2 mol)(15.999 g/mol) = 80.608 g

48.610 g Mg + 31.998 g O₂ → 80.608 g MgO

80.608 g reactants = 80.608 g products

- Challenge For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.
 - a. __Na(s) + __H₂O(l) → __NaOH(aq) + __H₂(g)

 $2Na(s) + 2H_sO(l) \rightarrow 2NaOH(aq) + 1H_s(q)$

2 atoms Na + 2 molecules H₂O --2 formula units NaOH + 1 molecule H₄

2 mol Na + 2 mol H₂O → 2 mol NaOH + 1 mol H₂

Maga:

2Na: (2 mol)(22.990 g/mol) = 45.980 g

2H₂O: (4 mol)(1.008 g/mol) + (2 mol)(15.999 g/mol) = 36.030 g

2NaOH: (2 mol)(22.990 g/mol) + (2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol) = 79.994 g

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 H_s : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g Na + 36.030 g H₂O → 79.994 g NaOH + 2.016 g H₂

82.01 g reactants = 82.01 g products

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