

Chemistry 2nd Semester Final Exam Review

Chemical Bonds

1. Give a physical description of how the atoms and molecules are arranged in solids, liquids, and gases.
A: In a liquid, the forces between the molecules are weak, so the molecules are free to move around. In a solid the forces between the molecules (intermolecular forces) are stronger so they are held in place.
2. Use the information from the previous question to describe the shape and volume for each phase (state) of matter.
*A: Solids have a repeating geometric pattern (lattice structure) in which the particles vibrate in place. They have a definite shape and a definite volume.
Particles within a liquid slide past each and are close, but are not held in place. They have a definite volume, but no definite shape.
Gases particles move chaotically bouncing off one another. They have neither a definite shape nor a definite volume.*
3. List the three phases of matter in order of increasing intermolecular attractions.
A: Gas, liquid, solid.
4. Why do the atoms and molecules in liquids move in a random pattern relative to one another instead of being in a solid form?
A: Their kinetic energy (temperature) is high enough to overcome the attractive forces that might otherwise hold them in a repeating pattern. The attractive forces aren't strong enough to hold them together at that temperature.

Conservation of Matter and Stoichiometry

5. The atomic mass unit and the mole are defined using what isotope? *A: Carbon-12*
6. Balance the following equations:
 - a. $\text{Al}_2\text{O}_3 + \text{Cl}_2 + \text{C} \rightarrow \text{AlCl}_3 + \text{CO}$
A: $\text{Al}_2\text{O}_3 + 3\text{Cl}_2 + 3\text{C} \rightarrow 2\text{AlCl}_3 + 3\text{CO}$
 - b. $\text{C}_4\text{H}_{10} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
A: $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$
 - c. $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$
A: $4\text{P} + 5\text{O}_2 \rightarrow 2\text{P}_2\text{O}_5$
7. What isotope is used to define one mole? *A: Carbon-12*
8. How many particles (atoms or molecules) are in one mole? *A: 6.02×10^{23}*
9. Find the molar mass for the following chemical formulas:
 - a. ZnF_2 *A: 103.377 g/mol*
 - b. $\text{Al}_2(\text{SO}_4)_3$ *A: 342.151 g/mol*
 - c. NH_4IO_3 *A: 192.941 g/mol*
10. How many moles are in 4.37 kg of NaOH? *A: 109.3 mol*
11. How many atoms are in 2075 g of He? *A: 3.12×10^{26} atoms*
12. What volume will 92.31 g of CO_2 occupy at STP? *A: 46.985 L*
13. How many grams of sodium hydroxide are needed to react with 6.23 g of barium bromide?
 $\text{NaOH} + \text{BaBr}_2 \rightarrow \text{NaBr} + \text{Ba(OH)}_2$
A: 1.68 g NaOH
14. In the reaction, $2\text{Mg} + \text{O}_2 \rightarrow 2\text{MgO}$, if 100.0 g of magnesium reacts with 50.0 g of oxygen, what mass of product is produced?
A: 126 g MgO

Gases and Their Properties

15. What is the kinetic molecular theory (KMT)?
A: The Kinetic Molecular Theory is used to explain the behavior of gases and is based upon the following postulates:
 - Gases are composed of a many particles that behave like hard spherical objects in a state of constant, random motion.
 - These particles move in a straight line until they collide with another particle or the walls of the container.
 - These particles are much smaller than the distance between particles, therefore the volume of a gas is mostly empty space and the volume of the gas molecule themselves is negligible.
 - There is no force of attraction between gas particles or between the particles and the walls of the container.
 - Collisions between gas particles or collisions with the walls of the container are elastic. That is, none of the energy of the gas particle is lost in a collision.

Chemistry Second Semester Final Exam Answers

Michael Seilmaier



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Biology Edward C. Gruber, 1963 Merck's Report Theodore Weicker, 1902 *Accounting and Auditing Examinations* Arco Publishing Company, 1949 *Bullying in a Small Town of China* Jinlong Wang, 101-01-01 In a small town of North China Fengjin generation after generation people have been suffering from the unfairness and injustice under an extreme social system The powerful and the rich have dominantly occupied core social resources The law of jungle prevails and tragedies exist everywhere Gongsun Ji was born in Fengjin the small town of North China From the moment he opened his eyes to the world he has experienced all kinds of unfairness and injustice that the small town of Fengjin brought to ordinary people as well as the disasters of people fighting and tearing each other apart The atmosphere of beating and scolding in Fengjin has a long history Gongsun Ji has no way of verifying where the tradition of beating and scolding in Fengjin originated from and why it started He only remembers the beating and scolding he experienced growing up and he doesn't know how long this vicious cycle of beating and scolding will continue To gain a foothold in the small town like Fengjin you have to be ruthless be tolerant or be gone In an extremely centralized and authoritarian society when social wealth is seriously unbalanced the powerful tends to hold power no matter how which will continue to strengthen the polarized distribution mechanism of social wealth leaving those innocent and helpless ordinary people who work hard without any hope of improvement If things go on like this it will inevitably arouse all kinds of social dissatisfaction and various social problems will arise in the micro areas Gongsun Ji did not know what the future society would be like but having grown up under an extreme system in a small town he had visited many parts of the world and seen some foreign civilizations under different social systems He once thought that some people and some things might be different The future belongs to the people of future but now is the present and people can still think more reflect and do something *General Test Guide for Civil Service Jobs* Arco Publishing Company, 1949 *Physics* Edward C. Gruber, 1963 *How to Pass Annapolis Entrance Exams* Arco Publishing Company, 1959 *Practice for the Army Tests* Arco Publishing Company, 1948 *From Coello to Inorganic Chemistry* Fred Basolo, 2012-12-06 From boyhood in the coal mining village of Coello Illinois to winning the Priestly Medal and becoming the president of the American Chemical Society Professor Emeritus Fred Basolo of Northwestern University traces the intertwined development of his life career and the field of inorganic chemistry With over a hundred photographs and dozens of structures and equations *From Coello to Inorganic Chemistry* details the major innovations travels family life and guests hosted while helping to build one of the world's leading inorganic chemistry departments from its humble beginnings at Northwestern University Students and chemists with interests in bioinorganic chemistry catalysis nanoscience new materials research and organometallics can follow the emergence of inorganic chemistry as a rival to organic chemistry through the accomplishments of one of its most influential pioneers **Implementation and Critical Assessment of the Flipped Classroom Experience** Scheg, Abigail G., 2015-01-31 In the past decade traditional classroom teaching models have been

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