

Identification of Ions Lab

Part 1: Pre-Lab Questions: Complete the pre-lab questions 1-5. You will use the solubility rules table for questions 5-6.

1. What makes water "hard"? (3pts)

The Magnesium and Calcium ions make water hard

2. After reviewing the prelab explain these terms in your own words: (2pt each)

Precipitate	A deposited solid from a solution
Alkaline Earth Elements	6 elements
Soluble	Has the ability to dissolve in water
Aqueous	Contains a lot of water or has water like properties

3. What happens when alkaline earth metals are combined with "hard" water? (4pts)

It'll Rust and dissolve in the water

4. Listed below are 2 reactants in a chemical reaction. Predict the products for these reactions. Follow example in a. (3 pt each)



$\text{NaCl} + \text{Pb}(\text{NO}_3)_2$		$\text{Na}(\text{NO}_3) + \text{PbCl}_2$
$\text{Na}_2\text{CO}_3 + \text{FeCl}_2$		$\text{Na}_2\text{Cl}_2 + \text{FeCO}_3$
$\text{BaBr}_2 + \text{Na}_2\text{SO}_4$		$\text{BaSO}_4 + \text{Na}_2\text{BR}_2$
$\text{AgNO}_3 + \text{MgI}_2$		$\text{AgI}_2 + \text{MgNO}_3$
$\text{Na}_3\text{PO}_4 + \text{Ni}(\text{ClO}_4)_2$		$\text{Na}_3(\text{ClO}_4)_2 + \text{NiPO}_4$

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