(i) Mole fraction: Commonly used symbol for mole fraction is x and subscript used on the right hand side of x denotes the component. It is defined as:

Mole fraction of a component = $\frac{\text{Number of moles of the component}}{\text{Number of moles of all the component}}$

(ii) Molality: Molality (m) is defined as the number of moles of the solute per kilogram (kg) of the solvent and is expressed as:

Molarity (m) = $\frac{\text{Moles of solute}}{\text{Mass of solvent in kg}}$

(iii) Molarity: Molarity (M) is defined as number of moles of solute dissolved in one litre (or one cubic decimetre) of solution.

Molarity $(M) = \frac{\text{Moles of solute}}{\text{Mass of solution in litre}}$

(iv) Mass percentage: The mass percentage of a component of a solution is defined as:

Mass % of a component = $\frac{\text{Mass of component in the solution}}{\text{Total mass of the solution}} \times 100$

- 4. Concentrated nitric acid used in laboratory work is 68% nitric acid by mass in aqueous solution. What should be the molarity of such a sample of the acid if the density of the solution is 1.504 g mL⁻¹?
- 4. Given:

Concentration of Nitric Acid, HNO₃ = 68%

Density of solution, d = 1.504 g/ml

To find: Molarity, Mo

Formula:

Density, $d = \frac{Mass[M]}{Volume[V]}$

Molarity, $M_s = \frac{\text{(Number of moles of solute)}}{\text{(Volume of solution in litres)}}$

68% of Nitric acid by mass in aqueous solution means that 68g [[68×100]/100] of Nitric acid present in 100g of solution.

 -63μ

- \Rightarrow Molecular mass of Nitric Acid, HNO₃ = [1 × 1] + [1 × 14] + [16 × 3]
- ⇒ Number of moles of Nitric Acid = [68/63]
 - = 1.079 moles
- ⇒ Given Density, d = 1.504 g/ml
- \Rightarrow Volume, v = [100/1.504]
 - = 66.489 ml
- ⇒ Molarity, Mo = [1.079/66.489] × 1000

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