

Binary Ionic Compounds

- Made up of _____ cations _____ anions.
- Name the _____ cation _____ anion _____.
- Name the _____ cation _____ anion _____, a charged ion, adding 1 to _____.
- Always include their ionic forms _____ that are fixed and _____ and (as is implied, not as _____ common monatomic _____ in parentheses after the name of the _____ metal in ionic form _____).

Examples: NaCl sodium chloride
 Cu_2S copper sulfide

AgCl magnesium chloride
 SnCl_2 tin (II) chloride

Binary Acid Compounds

- Made up of _____ cations _____ anions.
- Name the _____ that name the _____ polyatomic anion _____ will have changing the ending to "ide".

Examples: HNO_3 nitric acid
 HClO_4 perchloric acid

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CH Cherryholmes



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