Pages 163-164 Questions

1 molar mass of Fe₂(SO₄)₃ = (2 × 55.8) + 3 × [32.1 + (4 × 16.0)] = 399.9 g mol⁻¹ molar mass of Fe(OH)₃ = 55.8 + (3 × 17.0) = 106.8 g mol⁻¹ amount of Fe₂(SO₄)₃ (in moles) = mass/molar mass = 12.7 g/399.9 g mol⁻¹ = 0.03176 mol moles Fe(OH)₃ = moles Fe₂(SO₄)₃ × number of Fe(OH)₃ in equation/number of Fe₂(SO₄)₃ in equation = 0.03176 mol × 2 × 0.06352 mol mass of Fe(OH)₃ = moles × molar mass = 0.06352 mol × 106.8 g mol⁻¹ = 6.78 g

[e] It is a good idea to work out the malar masses of the two substances involved before starting the rest of the calculation.

2 molar mass of AgNO₃ = 107.9 + 14.0 + (3 × 16.0) = 169.9 g mol⁻¹ amount (in moles) of AgNO₃ = mass/molar mass = 12.6 g/169.9 g mol⁻¹ = 0.07416 mol ratio of Cu:AgNO₃ = 1:2 so, amount of copper = ½ × 0.07416 = 0.03708 mol mass of copper needed = moles × molar mass = 0.03708 mol × 63.5 g mol⁻¹ = 2.35 g

[e] The answer 2.35 g was obtained keeping all the numbers on the calculator during the calculation. If the rounded up value of 0.0371 mol is used, the answer 2.36 g is obtained. Either would score full marks.

A common error is to calculate the molar mass of $AgNO_1$ as 2×170 g mol⁻¹ because there are 2 moles of it in the equation. This is wrong. The number of moles of a substance depends only on its mass and not on the reaction. The stoichiometry comes into play when moles of one substance are converted to moles of another substance. This type of calculation can be done either by mass ratio or by converting to moles, then using the reaction stoichiometry and finally converting back to mass. The second method fits all types of calculation, so is the better one to use.

- molar mass of NaNO₃ = 23.0 + 14.0 + (3 × 16.0) = 85.0 g mol⁻¹ amount of NaNO₃ = mass/molar mass = 33.3 g/85.0 g mol⁻¹ = 0.392 mol ratio of O₃:NaNO₃ = 1:2 so, amount of oxygen = % × 0.392 = 0.196 mol volume of oxygen = moles × molar volume = 0.196 mol × 25.0 dm³ mol⁻¹ = 4.90 dm³
- 4 ratio of NH ::H = 2:3

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