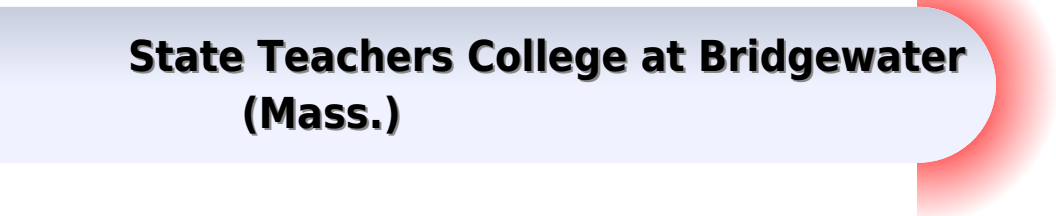


Chemistry First Nine Weeks study guide for exam

1. What are SI base units for time, length, amount of substance, mass?
2. Express numbers in scientific notation given standard, and given scientific put in standard.
 $452.7\text{ m} = ?$, $0.00077893\text{ L} = ?$, $6.77 \times 10^4\text{ mL} = ?$, $8.12 \times 10^{-6}\text{ km} = ?$
3. What are definitions of precision and accuracy and be able to apply them correctly to data given. See data tables in chapter 2 for examples.
4. Significant digits- how many in a given number, which types of zeroes are significant.
ex. 0.005603 m , 1241 m , 4.5800500 cm , 23.0 L
5. Correct rounding with significant digits. 4.5677 L to 3 sig. dig, 22.8655 to 4 sig.dig.
6. Adding, subtracting, multiplying, dividing correctly with sig. digits. (2 SEPARATE RULES)
 $34.41\text{ m} + 11.8764\text{ m} = ?$, $56.34\text{ m} - 21.9\text{ m} = ?$, $439\text{ cm} \times 5.33\text{ cm} = ?$, $189.22\text{ g} / 76\text{ m}$
7. Conversion of SI units to each other using dimensional analysis. You are responsible for the units you were given in class. Also make sure you know nano. Ex: $10^9\text{ nm} = 1\text{ m}$, very small unit, new branch of chemistry springing up around thee very tiny! You will hear more of this in future.
8. Density problems. Find the third term if given the other 2. Know the formula; it won't be given. Find density of a rock with mass of 35.8 g and volume of 8.1 mL . Find the volume of a substance with mass of 163 grams and density of 5.77 g/mL .
9. Dimensional analysis problems. See review problem sets for examples.
10. Characteristics of the 3 basic states of matter as they pertain to their container, volume, shape, etc.
11. Recognition of physical properties and chemical properties and/or changes. Copper carbonate which is green in color, when heated, turns black and loses mass. Grinding charcoal into a fine powder is an example of which?
12. What principles and means are used to separate mixtures?
13. Distinguish between homogeneous and heterogeneous mixtures. Well mixed koolaid, tossed salad, unsweetened iced tea, orange juice with pulp, solution.
14. Law of Conservation of Mass. State it and explain how you can use it to solve basic problems. Also % mass of an element in a compound. See practice problem sets.
15. State 4-5 parts of Dalton's atomic theory. Whose idea had he revived? Which part(s) do we not agree with completely anymore?
16. Three subatomic particles in an atom and where they are located. Be able to find all three of them if given atomic mass (and PT). Know how to write both hyphen notation and nuclear notation.
16. What were the characteristics of the nucleus that Rutherford discovered? What experiment did he discover them in?

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**State Teachers College at Bridgewater
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shows teachers how to implement AI tools in the classroom Developed for primary and secondary school teachers Learning with AI presents a powerful overview of the evolving trends of AI in education and offers invaluable insights into what artificial intelligence can accomplish in the classroom and beyond By learning how to use new AI tools and resources educators can empower themselves to navigate the challenges and seize the opportunities presented by AI From interactive learning techniques to advanced assignment and assessment strategies this comprehensive guide offers practical suggestions for integrating AI effectively into teaching and learning environments In the age of AI critical thinking skills and information literacy are more important than ever As AI continues to reshape the nature of human thinking and learning educators must develop and promote AI literacy to equip students with the skills they need to thrive in a rapidly evolving world This book serves as a compass guiding educators of all disciplines through the uncharted territory of AI powered education and the future of teaching and learning

Essentials of Physical Chemistry Don Shillady, 2011-07-27 At a time when U S high school students are producing low scores in mathematics and science on international examinations a thorough grounding in physical chemistry should not be considered optional for science undergraduates Based on the author s thirty years of teaching Essentials of Physical Chemistry merges coverage of calculus with chemistry and molecular physics in a friendly yet thorough manner Reflecting the latest ACS guidelines the book can be used as a one or two semester course and includes special topics suitable for senior projects The book begins with a math and physics review to ensure all students start on the same level and then discusses the basics of thermodynamics and kinetics with mathematics tuned to a level that stretches students abilities It then provides material for an optional second semester course that shows students how to apply their enhanced mathematical skills in a brief historical development of the quantum mechanics of molecules Emphasizing spectroscopy the text is built on a foundation of quantum chemistry and more mathematical detail and examples It contains sample classroom tested exams to gauge how well students know how to use relevant formulas and to display successful understanding of key concepts Coupling the development of mathematical skills with chemistry concepts encourages students to learn mathematical derivations Mini biographies of famous scientists make the presentation more interesting from a people point of view Stating the basic concepts of quantum chemistry in terms of analogies provides a pedagogically useful technique Covering key topics such as the critical point of a van der Waals gas the Michaelis Menten equation and the entropy of mixing this classroom tested text highlights applications across the range of chemistry forensic science pre medical science and chemical engineering In a presentation of fundamental topics held together by clearly established mathematical models the book supplies a quantitative discussion of the merged science of physical chemistry *Catalog*

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