



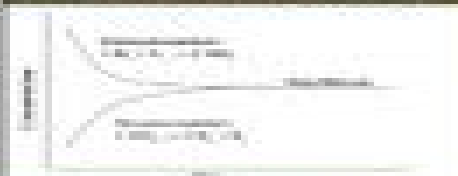
CHEMISTRY

CHEAT SHEET

EQUILIBRIUM

EQUILIBRIUM CHEAT SHEET

Chemical Equilibrium occurs when the concentrations of the chemical species remain the same with time (dynamic equilibrium). Equilibrium reactions are represented by the symbol $\rightleftharpoons$.

What is Equilibrium? Equilibrium is a state of balance between two or more opposing processes. In a chemical reaction, it is the state where the concentrations of the reactants and products remain constant over time. Equilibrium is a dynamic process, meaning that the reaction is still occurring, but the rates of the forward and reverse reactions are equal.	Equilibrium Constants K_c (Concentration): $K_c = \frac{[Products]}{[Reactants]}$ K_p (Pressure): $K_p = \frac{P_{Products}}{P_{Reactants}}$ K_a (Acid Dissociation): $K_a = \frac{[H^+][A^-]}{[HA]}$ K_b (Base Dissociation): $K_b = \frac{[OH^-][BH^+]}{[B]}$													
Equilibrium Constant, K_c For a general reaction: $aA + bB \rightleftharpoons cC + dD$ The equilibrium constant K_c is given by: $K_c = \frac{[C]^c [D]^d}{[A]^a [B]^b}$ Where $[A]$, $[B]$, $[C]$, and $[D]$ are the concentrations of the reactants and products at equilibrium.	Equilibrium Constant, K_p <table><tr><th>Reaction</th><th>Equilibrium Constant</th></tr><tr><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$</td><td>$K_p = \frac{P_{NH_3}^2}{P_{N_2} P_{H_2}^3}$</td></tr><tr><td>$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$</td><td>$K_p = \frac{P_{SO_3}^2}{P_{SO_2}^2 P_{O_2}}$</td></tr><tr><td>$PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$</td><td>$K_p = \frac{P_{PCl_3} P_{Cl_2}}{P_{PCl_5}}$</td></tr></table>		Reaction	Equilibrium Constant	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$	$K_p = \frac{P_{NH_3}^2}{P_{N_2} P_{H_2}^3}$	$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$	$K_p = \frac{P_{SO_3}^2}{P_{SO_2}^2 P_{O_2}}$	$PCl_5(g) \rightleftharpoons PCl_3(g) + Cl_2(g)$	$K_p = \frac{P_{PCl_3} P_{Cl_2}}{P_{PCl_5}}$				
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Equilibrium Constant, K_a For a weak acid HA in water: $HA \rightleftharpoons H^+ + A^-$ The acid dissociation constant K_a is given by: $K_a = \frac{[H^+][A^-]}{[HA]}$	Equilibrium Constant, K_b <table><tr><th>Reaction</th><th>Equilibrium Constant</th></tr><tr><td>$NH_3 + H_2O \rightleftharpoons NH_4^+ + OH^-$</td><td>$K_b = \frac{[NH_4^+][OH^-]}{[NH_3]}$</td></tr><tr><td>$CH_3COO^- + H_2O \rightleftharpoons CH_3COOH + OH^-$</td><td>$K_b = \frac{[CH_3COOH][OH^-]}{[CH_3COO^-]}$</td></tr><tr><td>$CO_3^{2-} + H_2O \rightleftharpoons HCO_3^- + OH^-$</td><td>$K_b = \frac{[HCO_3^-][OH^-]}{[CO_3^{2-}]}$</td></tr></table>		Reaction	Equilibrium Constant	$NH_3 + H_2O \rightleftharpoons NH_4^+ + OH^-$	$K_b = \frac{[NH_4^+][OH^-]}{[NH_3]}$	$CH_3COO^- + H_2O \rightleftharpoons CH_3COOH + OH^-$	$K_b = \frac{[CH_3COOH][OH^-]}{[CH_3COO^-]}$	$CO_3^{2-} + H_2O \rightleftharpoons HCO_3^- + OH^-$	$K_b = \frac{[HCO_3^-][OH^-]}{[CO_3^{2-}]}$				
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Le Chatelier's Principle: The effect of changing conditions on equilibrium <table><tr><th>Condition</th><th>Effect on Equilibrium</th><th>Example</th></tr><tr><td>Concentration</td><td> - Increasing the concentration of a reactant shifts the equilibrium to the right. - Increasing the concentration of a product shifts the equilibrium to the left. </td><td>$CO(g) + 2H_2(g) \rightleftharpoons CH_3OH(g)$ Adding more CO or H_2 increases the yield of CH_3OH.</td></tr><tr><td>Pressure</td><td> - Increasing the pressure shifts the equilibrium to the side with fewer moles of gas. - Decreasing the pressure shifts the equilibrium to the side with more moles of gas. </td><td>$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ Increasing pressure increases the yield of NH_3.</td></tr><tr><td>Temperature</td><td> - For an exothermic reaction, increasing temperature shifts the equilibrium to the left. - For an endothermic reaction, increasing temperature shifts the equilibrium to the right. </td><td>$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$ (exothermic) Increasing temperature decreases the yield of SO_3.</td></tr></table>			Condition	Effect on Equilibrium	Example	Concentration	- Increasing the concentration of a reactant shifts the equilibrium to the right. - Increasing the concentration of a product shifts the equilibrium to the left.	$CO(g) + 2H_2(g) \rightleftharpoons CH_3OH(g)$ Adding more CO or H_2 increases the yield of CH_3OH .	Pressure	- Increasing the pressure shifts the equilibrium to the side with fewer moles of gas. - Decreasing the pressure shifts the equilibrium to the side with more moles of gas.	$N_2(g) + 3H_2(g) \rightleftharpoons 2NH_3(g)$ Increasing pressure increases the yield of NH_3 .	Temperature	- For an exothermic reaction, increasing temperature shifts the equilibrium to the left. - For an endothermic reaction, increasing temperature shifts the equilibrium to the right.	$2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$ (exothermic) Increasing temperature decreases the yield of SO_3 .
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Chemistry Equilibrium Study Guide Answers

David R. Klein



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O.

Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter *Chemistry, Student Study Guide*

James E. Brady, Fred Senese, 2008-01-28 The image on the front cover depicts a carbon nanotube emerging from a glowing plasma of hydrogen and carbon as it forms around particles of a metal catalyst Carbon nanotubes are a recently discovered allotrope of carbon Three other allotropes of carbon buckyballs graphite and diamond are illustrated at the left as is the molecule methane CH_4 from which nanotubes and buckyballs can be made The element carbon forms an amazing number of compounds with structures that follow from simple methane found in natural gas to the complex macromolecules that serve as the basis of life on our planet The study of chemistry also follows from the simple to the more complex and the strength of this text is that it enables students with varied backgrounds to proceed together to significant levels of achievement

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Organic Chemistry, Student Study Guide and Solutions Manual David R. Klein, 2017-01-04 This is the Student Study Guide and Solutions Manual to accompany Organic Chemistry 3e Organic Chemistry 3rd Edition is not merely a compilation of principles but rather it is a disciplined method of thought and analysis Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems Readers must learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry Existing textbooks provide extensive coverage of the principles but

there is far less emphasis on the skills needed to actually solve problems **Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application *Organic Chemistry, 5e Student Study Guide and Solutions Manual* David R. Klein, Laurie S.

Starkey, 2025-03-18 Success in organic chemistry requires mastery in two core aspects fundamental concepts and the skills needed to apply those concepts and solve problems With Organic Chemistry Student Study Guide and Solutions Manual 5th Edition students can learn to become proficient at approaching new situations methodically based on a repertoire of skills These skills are vital for successful problem solving in organic chemistry

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Answers to Even- Numbered Exercises 9. Experiment with the xman utility to answer the following questions: a. How many man pages are in the Devices section of the manual? Answers to Odd-Numbered Problems CHAPTER 1. Exercises 1.1. 1. (a) ordinary, first order. (c) partial, second order. (e) ordinary, third order. (g) ordinary, second order. Answers to Even-Numbered Exercises How can you keep other users from using write to communicate with you? Why would you want to? Give the command mesg n to keep ordinary users from writing to ... Why do some science or math books only have answers ... Jan 30, 2015 — Some science and math books only provide answers to odd or even numbered questions as a way to encourage students to practice ... MARK G. SOBELL A PRACTICAL GUIDE TO LINUX ... by MG SOBELL · 2013 · Cited by 55 — ... EXERCISES. 1. The following message is displayed when you attempt to log in with an incorrect username or an incorrect password: Page 81. ADVANCED EXERCISES ... ANSWERS TO EVEN-NUMBERED EXERCISES - Sobell Jul 27, 2013 — Answers to Even-numbered Exercises
. 1. Wile?
. 2. What does the /etc/resolv.conf file do? What do the nameserver lines in
. 1 Answers to Chapter 3, Odd-numbered Exercises 1 Answers to Chapter 3, Odd-numbered Exercises. 1) $r(n) = 25r(n - 1) + 3r(n - 2) + 10n - 1$. There are $25r(n - 1)$ identifiers satisfying the first condition, $3r$... Vim Question - Single command to swap words Jan 5, 2012 — Hi, I'm working through Sobell's book Linux Commands, Editors and Shell ... odd-numbered exercises (for which he does not publish the answers). Why do textbooks often include the solutions to odd or ... Jun 18, 2019 — My question is, why do textbooks often include the solutions to odd or even numbered problems but not both? In my case, I don't think space is ... Directed Reading A Holt Science and Technology. 4. The Properties of Matter. Section: Physical ... Answer Key. TEACHER RESOURCE PAGE. Page 5. 31. Answers will vary. Sample answer ... Chemical Properties Answer.pdf A matter with different properties is known as a(n) a. chemical change. b. physical change. c. chemical property. d. physical property. Directed Reading A 3. A substance that contains only one type of particle is a(n). Pure Substance ... Holt Science and Technolnov. 4. Elements. Compounds, and Mixtures. Page 5. Name. Directed Reading Chapter 3 Section 3 . Holt Science and Technology. 5. Minerals of the Earth's Crust. Skills Worksheet. Directed

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