

General Chemistry II Chapter 15 Study Guide

- $S'_{\text{gas}} > S'_{\text{liquid}} > S'_{\text{solid}}$
- For similar molecules, S' tends to increase with:
 - Increasing molar mass
 - Increased number of atoms in the formula

Standard entropy of a reaction ($\Delta S'_{\text{rxn}}$):

- $\Delta S'_{\text{rxn}}$: is the entropy change that occurs when all reactants and products are in their standard states:

$$\Delta S'_{\text{rxn}} = \sum m S'_{\text{products}} - \sum n S'_{\text{reactants}}$$

where m and n are the mols of products and reactants, given by the coefficients in the balanced equation.

Suppose a reaction: $a A + b B \rightarrow c C + d D$

$$\sum m S'_{\text{products}} = c \Delta S' (C) + d \Delta S' (D)$$

$$\sum n S'_{\text{reactants}} = a \Delta S' (A) + b \Delta S' (B)$$

Remember to multiply $\Delta H'$ of each substance by its coefficient.

Predicting the sign of $\Delta S'_{\text{rxn}}$:

- We can often predict the sign of $\Delta S'_{\text{rxn}}$ for processes that involve a change in the number of moles of gas
 - $\Delta S'_{\text{rxn}}$ is **positive** if the amount of gas increases
 - $\Delta S'_{\text{rxn}}$ is **negative** if the amount of gas decreases
 - $\Delta S'_{\text{rxn}}$ is **likely to be positive** if a new structure forms that has more freedom of motion

Gibbs free energy:

- Gibbs free energy (G): combines the enthalpy and entropy of a system:
 - $G = H - TS$
- Gibbs free energy change (ΔG): is a measure of the spontaneity of a process and of the useful energy available from it:
 - $\Delta G_{\text{rxn}} = \Delta H_{\text{rxn}} - T\Delta S_{\text{rxn}}$
- $\Delta G < 0$ for a spontaneous reaction ($-\Delta G$)
- $\Delta G > 0$ for a nonspontaneous reaction ($+\Delta G$)
- $\Delta G = 0$ for a process at equilibrium ($\Delta G = 0$)

Chemistry Chapter 15 Study Guide Answers

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