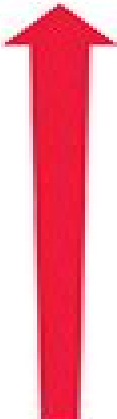
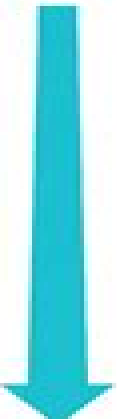


# Acids and Bases

| Acid                                                                               |                       |                          |                                                            | Base                                                       |                          |                                                                                      |  |
|------------------------------------------------------------------------------------|-----------------------|--------------------------|------------------------------------------------------------|------------------------------------------------------------|--------------------------|--------------------------------------------------------------------------------------|--|
| Strength of Acid                                                                   | K <sub>a</sub>        | Name                     | Formula                                                    | Formula                                                    | Name                     | Strength of Base                                                                     |  |
| <b>Strongest</b>                                                                   |                       |                          |                                                            |                                                            |                          | <b>Weakest</b>                                                                       |  |
|   | Large                 | Perchloric acid          | HClO <sub>4</sub>                                          | ClO <sub>4</sub> <sup>-</sup>                              | Perchlorate ion          |  |  |
|                                                                                    | 3.2×10 <sup>8</sup>   | Hydroiodic acid          |                                                            | I <sup>-</sup>                                             | Iodide                   |                                                                                      |  |
|                                                                                    | 1.0×10 <sup>9</sup>   | Hydrobromic acid         | HBr                                                        | Br <sup>-</sup>                                            | Bromide                  |                                                                                      |  |
|                                                                                    | 1.3×10 <sup>9</sup>   | Hydrochloric acid        | HCl                                                        | Cl <sup>-</sup>                                            | Chloride                 |                                                                                      |  |
|                                                                                    | 1.0×10 <sup>3</sup>   | Sulfuric acid            | H <sub>2</sub> SO <sub>4</sub>                             | HSO <sub>4</sub> <sup>-</sup>                              | Hydrogen sulfate ion     |                                                                                      |  |
|                                                                                    | 2.4×10 <sup>1</sup>   | Nitric acid              | HNO <sub>3</sub>                                           | NO <sub>3</sub> <sup>-</sup>                               | Nitrate ion              |                                                                                      |  |
|                                                                                    | -----                 | Hydronium ion            | H <sub>3</sub> O <sup>+</sup>                              | H <sub>2</sub> O                                           | Water                    |                                                                                      |  |
|                                                                                    | 5.4×10 <sup>-2</sup>  | Oxalic acid              | HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> H            | HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>-</sup> | Hydrogen oxalate ion     |                                                                                      |  |
|                                                                                    | 1.3×10 <sup>-2</sup>  | Sulfurous acid           | H <sub>2</sub> SO <sub>3</sub>                             | HSO <sub>3</sub> <sup>-</sup>                              | Hydrogen sulfite ion     |                                                                                      |  |
|                                                                                    | 1.0×10 <sup>-2</sup>  | Hydrogen sulfate ion     | HSO <sub>4</sub> <sup>-</sup>                              | SO <sub>4</sub> <sup>2-</sup>                              | Sulfate ion              |                                                                                      |  |
|                                                                                    | 7.1×10 <sup>-3</sup>  | Phosphoric acid          | H <sub>3</sub> PO <sub>4</sub>                             | H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>                | Dihydrogen phosphate ion |                                                                                      |  |
|                                                                                    | 7.2×10 <sup>-4</sup>  | Nitrous acid             | HNO <sub>2</sub>                                           | NO <sub>2</sub> <sup>-</sup>                               | Nitrite ion              |                                                                                      |  |
|                                                                                    | 6.6×10 <sup>-4</sup>  | Hydrofluoric acid        | HF                                                         | F <sup>-</sup>                                             | Fluoride ion             |                                                                                      |  |
|                                                                                    | 1.8×10 <sup>-4</sup>  | Methanoic acid           | HCO <sub>2</sub> H                                         | HCO <sub>2</sub> <sup>-</sup>                              | Methanoate ion           |                                                                                      |  |
|                                                                                    | 6.3×10 <sup>-5</sup>  | Benzoic acid             | C <sub>6</sub> H <sub>5</sub> COOH                         | C <sub>6</sub> H <sub>5</sub> COO <sup>-</sup>             | Benzoate ion             |                                                                                      |  |
|                                                                                    | 5.4×10 <sup>-5</sup>  | Hydrogen oxalate ion     | HO <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>-</sup> | O <sub>2</sub> C <sub>2</sub> O <sub>2</sub> <sup>2-</sup> | Oxalate ion              |                                                                                      |  |
|                                                                                    | 1.8×10 <sup>-5</sup>  | Ethanoic acid            | CH <sub>3</sub> COOH                                       | CH <sub>3</sub> COO <sup>-</sup>                           | Ethanoate (acetate) ion  |                                                                                      |  |
|                                                                                    | 4.4×10 <sup>-7</sup>  | Carbonic acid            | H <sub>2</sub> CO <sub>3</sub>                             | HCO <sub>3</sub> <sup>-</sup>                              | Hydrogen carbonate ion   |                                                                                      |  |
|                                                                                    | 1.1×10 <sup>-7</sup>  | Hydrosulfuric acid       | H <sub>2</sub> S                                           | HS <sup>-</sup>                                            | Hydrogen sulfide ion     |                                                                                      |  |
|                                                                                    | 6.3×10 <sup>-8</sup>  | Dihydrogen phosphate ion | H <sub>2</sub> PO <sub>4</sub> <sup>-</sup>                | HPO <sub>4</sub> <sup>2-</sup>                             | Hydrogen phosphate ion   |                                                                                      |  |
|  | 6.2×10 <sup>-8</sup>  | Hydrogen sulfite ion     | HS <sup>-</sup>                                            | SO <sub>3</sub> <sup>2-</sup>                              | Sulfite ion              |  |  |
|                                                                                    | 2.9×10 <sup>-8</sup>  | Hypochlorous acid        | HClO                                                       | ClO <sup>-</sup>                                           | Hypochlorite ion         |                                                                                      |  |
|                                                                                    | 6.2×10 <sup>-10</sup> | Hydrocyanic acid         | HCN                                                        | CN <sup>-</sup>                                            | Cyanide ion              |                                                                                      |  |
|                                                                                    | 5.8×10 <sup>-10</sup> | Ammonium ion             | NH <sub>4</sub> <sup>+</sup>                               | NH <sub>3</sub>                                            | Ammonia                  |                                                                                      |  |
|                                                                                    | 5.8×10 <sup>-10</sup> | Boric acid               | H <sub>3</sub> BO <sub>3</sub>                             | H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>                | Dihydrogen borate        |                                                                                      |  |
|                                                                                    | 4.7×10 <sup>-11</sup> | Hydrogen carbonate ion   | HCO <sub>3</sub> <sup>-</sup>                              | CO <sub>3</sub> <sup>2-</sup>                              | Carbonate ion            |                                                                                      |  |
|                                                                                    | 4.2×10 <sup>-13</sup> | Hydrogen phosphate ion   | HPO <sub>4</sub> <sup>2-</sup>                             | PO <sub>4</sub> <sup>3-</sup>                              | Phosphate ion            |                                                                                      |  |
|                                                                                    | 1.8×10 <sup>-13</sup> | Dihydrogen borate ion    | H <sub>2</sub> BO <sub>3</sub> <sup>-</sup>                | HBO <sub>3</sub> <sup>2-</sup>                             | Hydrogen borate ion      |                                                                                      |  |
|                                                                                    | 1.3×10 <sup>-13</sup> | Hydrogen sulfide ion     | HS <sup>-</sup>                                            | S <sup>2-</sup>                                            | Sulfide ion              |                                                                                      |  |
|                                                                                    | 1.6×10 <sup>-14</sup> | Hydrogen borate ion      | HBO <sub>3</sub> <sup>2-</sup>                             | BO <sub>3</sub> <sup>3-</sup>                              | Borate ion               |                                                                                      |  |
|                                                                                    | -----                 | water                    | H <sub>2</sub> O                                           | OH <sup>-</sup>                                            | Hydroxide                |                                                                                      |  |
| <b>Weakest</b>                                                                     |                       |                          |                                                            |                                                            |                          | <b>Strongest</b>                                                                     |  |

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Morgan, George Edwin Rines, E. T.  
Roe, Nathan Haskell Dole, Edward  
Thomas Roe, Thomas Campbell  
Copeland**

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