

When stoichiometric equivalent moles of acid and base react, they are said to be at the *equivalence point*. The process of reaching the equivalence point with acids and bases is called *neutralization*.

In order to know when the equivalence point is reached, an indicator is added to the titration vessel. An *indicator* is a compound that will react with one of the titration species and for which a color change is observed very near the equivalence point, when the appropriate indicator has been selected. The indicator used in this experiment is called phenolphthalein which turns from clear in acid solution to a pink color when excess base is present.

Sample calculation

If 0.558g of potassium hydrogen phthalate (KHP) is dissolved in 50 mL of deionized water in a flask, what is the concentration of a NaOH solution if 12.78 mL are required to titrate the potassium hydrogen phthalate to the equivalence point?

- a) Determine the moles of potassium hydrogen phthalate (KHP) in the flask using mass and molar mass:

$$\text{moles KHP} = \text{mass KHP} \times \frac{1}{M_{\text{KHP}}} = 0.558\text{g} \times \frac{1 \text{ mol}_{\text{KHP}}}{204.23\text{g}_{\text{KHP}}} = 0.00273 \text{ mol}_{\text{KHP}}$$

- b) Using the stoichiometric relationship in the balanced chemical reaction (Eq. 1), determine the moles of base required to reach the equivalence point:

$$\text{mole NaOH} = \text{mol}_{\text{KHP}} \times \frac{1 \text{ mol}_{\text{NaOH}}}{1 \text{ mol}_{\text{KHP}}} = 0.00273 \text{ mol}_{\text{KHP}} \times \frac{1 \text{ mol}_{\text{NaOH}}}{1 \text{ mol}_{\text{KHP}}} = 0.00273 \text{ mol}_{\text{NaOH}}$$

- c) Determine the concentration of NaOH by dividing the number of mole by volume of NaOH solution added:

$$\text{molarity of NaOH} = \frac{0.00273 \text{ mol}_{\text{NaOH}}}{0.01278 \text{ L}_{\text{NaOH solution}}} = 0.214 \text{ M NaOH}$$

After the $\approx 0.3 \text{ M}$ NaOH solution has been standardized (its accurate average molarity determined by titration), it can be referred to as a standardized solution and its average molarity from the three titrations should be written on the bottle. It, in turn, will be employed to determine the acetic acid content in household vinegar. Like potassium hydrogen phthalate, acetic acid (ethanoic acid) has one acidic proton, H^+ . Because the acid is monoprotic, it will react one-to-one with NaOH (Eq. 2).



This time you will be starting with a dilute solution of acetic acid in the flask instead of a solid, and you will know the accurate molarity of the NaOH solution (the average of the three values determined). After the titration, knowing the volume of NaOH used in the titration, you can determine the mole of NaOH needed to reach the equivalence point with the acetic acid:

$$\text{moles NaOH} = M_{\text{NaOH}} \left(\frac{\text{mol}}{\text{L}} \right) \times V_{\text{NaOH}} (\text{L})$$

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