

Boiling Curve

The following is a plot showing typical behaviour in a liquid when it is heated while the pressure is atmospheric. Generally, temperature increases as energy is added until it reaches a point at which it is no longer possible to increase its temperature. This point is called the boiling point. At a given temperature, it is possible to observe that both water and steam exist at a given temperature. It is important to realise that both water and steam exist at the same time at a given temperature in equilibrium. The amount of energy that is needed to heat a substance up to its boiling point is called the **specific latent heat**. It is usually denoted by L_v . The temperature of the substance does not change during the boiling process. A graph of temperature against energy added is shown in Figure 1. The graph shows a horizontal line at the boiling point. A point on the horizontal line is marked as the boiling point of water.



Answer the questions below by referring to Figure 1.

- The temperature of the water at the boiling point is 100°C.
- The temperature of the steam at the boiling point is 100°C.
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- The amount of energy added to the water to reach the boiling point is 100 J.
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