

balance nuclear reaction equations

aluminum-26 undergoes positron emission
potassium-40 undergoes beta decay (electron)
uranium-238 undergoes alpha decay



electron configurations

Write the electron configuration for C.

Write the electron configuration for Cr.

Write the shorthand notation for Bi.



NOMENCLATURE (naming compounds, writing formulas)

Write the formula for each of the following compounds:

- Sulfur hexafluoride SF_6
- Lithium nitride Li_3N
- Chromium (III) carbonate $\text{Cr}_2(\text{CO}_3)_3$
- Tin (II) chloride SnCl_2
- Ammonium acetate $\text{NH}_4\text{C}_2\text{H}_3\text{O}_2$ (or $\text{CH}_3\text{COONH}_4$)
- Mercury (I) chloride Hg_2Cl_2
- Potassium bromate KBrO_3
- Hydrosulfuric acid $\text{H}_2\text{S(aq)}$
- Chloric acid HClO_3
- Sulfurous acid H_2SO_3

Name each of the following compounds:

- CuSO_4 Copper (II) sulfate
- AlF_3 aluminum fluoride
- HI hydrogen iodide
- NO nitrogen monoxide
- $\text{H}_2\text{Se(aq)}$ hydroselenic acid
- HNO_3 nitric acid
- HNO_2 nitrous acid
- NaHSO_4 sodium hydrogen sulfate

Half-life

Tritium (H-3) is a radioactive isotope of hydrogen with a half-life of 12.3 years. How long would it take for a 40.0 g sample to decay down to 1.25 g?

$$40 \rightarrow 20 \rightarrow 10 \rightarrow 5 \rightarrow 2.5 \rightarrow 1.25$$

1 2 3 4 5

61.5 yr

Fe-61 has a half-life of 6.00 min. Of a 100.0 mg sample, how much will remain after 18.0 min?

$$3 \text{ half lives} \quad 100 \rightarrow 50 \rightarrow 25 \rightarrow 12.5 \text{ mg}$$

1 2 3

Average atomic mass

The element copper has naturally occurring isotopes with mass numbers of 63 and 65.

The relative abundance and atomic masses are 69.2% for a mass of 63amu and 30.8% for a mass of 65amu. Calculate the average atomic mass of copper.

$$\begin{array}{rcl} 63 \times .692 & = & 43.596 \rightarrow 44 \\ 65 \times .308 & = & 20.02 \rightarrow 20. \\ + & & \\ \hline & & 63.62 \rightarrow 64 \end{array}$$

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Puja Mehta



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 - The Rise of Digital Reading Chemistry Extra Final Exam Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Extra Final Exam Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Extra Final Exam Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Extra Final Exam Study Guide Answers
 - Personalized Recommendations
 - Chemistry Extra Final Exam Study Guide Answers User Reviews and Ratings
 - Chemistry Extra Final Exam Study Guide Answers and Bestseller Lists

5. Accessing Chemistry Extra Final Exam Study Guide Answers Free and Paid eBooks
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 - Highlighting and Note-Taking Chemistry Extra Final Exam Study Guide Answers
 - Interactive Elements Chemistry Extra Final Exam Study Guide Answers
8. Staying Engaged with Chemistry Extra Final Exam Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Extra Final Exam Study Guide Answers
9. Balancing eBooks and Physical Books Chemistry Extra Final Exam Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Extra Final Exam Study Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Extra Final Exam Study Guide Answers
 - Setting Reading Goals Chemistry Extra Final Exam Study Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Extra Final Exam Study Guide Answers
 - Fact-Checking eBook Content of Chemistry Extra Final Exam Study Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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