

General Chemistry II Chapter 14 Study Guide

Types of reaction rates:

- If the concentration of a reactant is plotted against time:
 - If the graph has a curved line = rate changes; will reach equilibrium
 - If the graph is linear = rate is constant; will run out of reactant
- Average rate of reaction: change in concentration of a reactant or product over time; reaction rate = slope
- instantaneous rate: slope of a line tangent to the curve; can pick any time
- Initial rate: rate within $t = 0$ to $t =$ given time; generally portion of the curve that is linear

Rate laws:

- Rate laws: only reactants appear in the rate law; the exponents tells the order of the rxn with respect to each reactant $\rightarrow$ sum of exponents is the overall order of the rxn; once we know the order of each reactant, we can plug in the concentrations and reaction rate to calculate the rate constant, k
- rate = $k[A]^m[B]^n$

Determining reaction orders:

- Suppose A and B are two reactants in a reaction:
 - If double $[A]$ will cause no change to rate, the order on A is 0 ($m=0$), e.g. $[2]0 = 1$, $[4]0 = 1$
 - If double $[A]$ will cause the rate to double, the order on A is 1 ($m=1$), e.g. $[2]1 = 2$, $[4]1 = 4$
 - If double $[A]$ will cause the rate to quadruple, the order on A is 2 ($m=2$), e.g. $[2]2 = 4$, $[4]2 = 16$
 - The order on reactant is not related to its coefficient in the chemical equation.
 - the exponents must be determined experimentally.

Determining rate laws: isolation method

- For rxns with more than one reactant, the rate law can be determined by the isolation method; approach:
 1. Fix the concentration of all reactants except for one $\rightarrow$ determine the rxn rates with varied concentrations of that reactant
 2. The change in rate is only attributed to that one reactant, enabling determination of the order of that reactant in the rate law
 3. Apply similar method to determine the order of other reactants

Zero-order process: overall order = 0:

Zero-Order Process: overall order = 0

For reaction $A \rightarrow \text{Product}$:

$$\text{Rate} = k[A]^0 = k$$

$$-\frac{d[A]}{dt} = k$$

$$d[A] = -k dt$$

Take the integral:

$$[A]_t = -kt + C$$

Plug in initial condition $[A] = [A]_0$ at $t=0$:

$$C = [A]_0$$

Units of k : $k = \text{rate} = \text{M/s}$



Zero-Order Process

Plot of $[A]_t$ vs. t



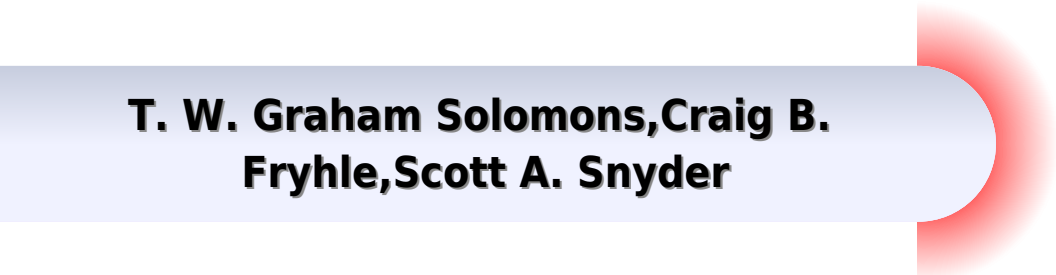
$$[A]_t = -kt + [A]_0$$

Equivalent to a straight-line function:
 $y = mx + b$

- Plot of $[A]_t$ vs. t is a straight line with $-k$ as slope.
- b is the y-intercept (i.e. the initial concentration $[A]_0$)
- Eventually $[A]$ will go to zero (all reactants will be used up)

Chemistry Chapter 14 Study Guide Answers

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Fryhle, Scott A. Snyder**



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