

**Electrochemistry Practice Test**

Name:

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**I. Multiple Choice**

1. When a metal undergoes corrosion, it:

- (a) loses electrons  
(b) becomes reduced

- (c) acts as an oxidizing agent  
(d) decreases in oxidation number

2. At standard conditions,  $\text{Fe}^{+2}$  reacts spontaneously with

- (a)  $\text{I}_2$   
(b) Co

- (c)  $\text{Br}^-$   
(d)  $\text{Ag}^+$

3. Which of the following half-reactions is balanced?

- (a)  $\text{SO}_4^{-2} + \text{H}_2\text{O} \rightleftharpoons \text{SO}_3^{-2} + 2\text{H}^+ + 2\text{e}^-$   
(b)  $\text{SO}_4^{-2} + \text{H}_2\text{O} + 2\text{e}^- \rightleftharpoons \text{SO}_3^{-2} + 2\text{H}^+$   
(c)  $\text{SO}_4^{-2} + 2\text{H}^+ + 2\text{e}^- \rightleftharpoons \text{SO}_3^{-2} + \text{H}_2\text{O}$   
(d)  $\text{SO}_4^{-2} + 2\text{H}^+ \rightleftharpoons \text{SO}_3^{-2} + \text{H}_2\text{O} + 2\text{e}^-$

4. During a redox reaction, the oxidizing agent:

- (a) reduces other species  
(b) gains electrons

- (c) increases in oxidation number  
(d) becomes oxidized

5. For a given redox rxn, the oxidation # of tin changed from +2 to +4. As a result, tin:

- (a) lost 2 electrons and was reduced  
(b) gained 2 electrons and was reduced  
(c) lost 2 electrons and was oxidized  
(d) gained 2 electrons and was oxidized

6. In which of the following compounds does carbon have an oxidation number of -2?

- (a) CO  
(b)  $\text{CO}_2$

- (c)  $\text{CH}_2\text{O}$   
(d)  $\text{CH}_3\text{OH}$

7. Which of the following equations represents a redox reaction?

- (a)  $\text{ZnCl}_2(\text{aq}) \rightleftharpoons \text{Zn}^{+2}(\text{aq}) + 2\text{Cl}^-(\text{aq})$   
(b)  $\text{Zn}(\text{s}) + \text{Br}_2(\text{l}) \rightleftharpoons \text{ZnBr}_2(\text{s})$   
(c)  $\text{H}_2\text{CO}_3(\text{aq}) \rightleftharpoons \text{H}_2\text{O}(\text{l}) + \text{CO}_2(\text{g})$   
(d)  $2\text{NaI}(\text{aq}) + \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightleftharpoons \text{PbI}_2(\text{s}) + 2\text{NaNO}_3(\text{aq})$

8. Consider the following reaction:

# Chemistry 12 Electrochemistry Test

**Peter Schwamb, Allyne Litchfield  
Merrill**

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