

% Composition and Stoichiometry Test SCH3U Version #4

/18K /18T /5C /4A

MULTIPLE CHOICE: Choose the most correct answer. (10 P)

1. A possible molecular formula for the empirical formula $C_2H_4PNO_4$ is?

- a. $C_4H_8PNO_8$ b. $C_2H_4N_2O_4$ c. $C_2H_4PNO_4$ d. $C_2H_4O_4$ e. none of the above

2. Compound X has a molar mass of 154.18 g/mol and an empirical formula of C_6H_8N . What is the molecular formula of this compound?

- a. C_6H_8N b. $C_{12}H_{16}N_2$ c. $C_{18}H_{24}N_3$ d. $C_{12}H_{16}N$ e. $C_{18}H_{24}N$

3. The total number of hydrogen atoms in 6.7 mol of ammonium hydroxide (NH_4OH) is:

- a. 4.02×10^{24} b. 2.42×10^{24} c. 1.81×10^{24} d. 2.00×10^{24} e. 4.00×10^{24}

4. The number of moles of oxygen in 2 moles of copper (II) sulfate pentahydrate, $CuSO_4 \cdot 5H_2O$ is?

- a. 10 b. 14 c. 18 d. 8 e. 2

5. What is the mass percent of water in $Na_2CO_3 \cdot 4H_2O$?

- a. 41 % b. 34 % c. 50 % d. 57 % e. 28 %

6. The number of molecules of hydrogen peroxide (H_2O_2) found in 2.75 mol is?

- a. 1.65×10^{24} b. 4.97×10^{24} c. 4.97×10^{23} d. 1.65×10^{23} e. 4.97×10^{23}

7. The limiting reagent of a chemical reaction is:

- a. the reactant of a chemical process that is not consumed completely
b. the product of a chemical process that is consumed completely
c. the product of a chemical process that is not consumed completely
d. the reactant of a chemical process that is consumed completely
e. none of the above

8. The combustion of propane (C_3H_8) in the presence of excess oxygen yields CO_2 and H_2O .

$$C_3H_8(g) + 5O_2(g) \rightarrow 3CO_2(g) + 4H_2O(g)$$

When 2.5 mol of O_2 are consumed in this reaction, _____ mol of CO_2 are produced.

- A) 1.5 B) 3.0 C) 5.0 D) 6.0 E) 2.5

9. The molar mass of copper(II) sulfate pentahydrate, $CuSO_4 \cdot 5H_2O$, is

- a. 245.6 g/mol b. 177.5 g/mol c. 219.2 g/mol d. 241.5 g/mol e. 249.6 g/mol

10. What is the mass in grams of $AlCl_3$ for a sample that contains 5.44×10^{23} chloride ions?

- a. 1 g b. 2 g c. 3 g d. 4 g e. 5 g

Chemistry 11 Stoichiometry Test

Daniel F McAuley



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1. Understanding the eBook Chemistry 11 Stoichiometry Test
 - The Rise of Digital Reading Chemistry 11 Stoichiometry Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry 11 Stoichiometry Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry 11 Stoichiometry Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry 11 Stoichiometry Test
 - Personalized Recommendations
 - Chemistry 11 Stoichiometry Test User Reviews and Ratings
 - Chemistry 11 Stoichiometry Test and Bestseller Lists
5. Accessing Chemistry 11 Stoichiometry Test Free and Paid eBooks
 - Chemistry 11 Stoichiometry Test Public Domain eBooks
 - Chemistry 11 Stoichiometry Test eBook Subscription Services
 - Chemistry 11 Stoichiometry Test Budget-Friendly Options
6. Navigating Chemistry 11 Stoichiometry Test eBook Formats

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- Chemistry 11 Stoichiometry Test Compatibility with Devices
- Chemistry 11 Stoichiometry Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry 11 Stoichiometry Test
 - Highlighting and Note-Taking Chemistry 11 Stoichiometry Test
 - Interactive Elements Chemistry 11 Stoichiometry Test
- 8. Staying Engaged with Chemistry 11 Stoichiometry Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry 11 Stoichiometry Test
- 9. Balancing eBooks and Physical Books Chemistry 11 Stoichiometry Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry 11 Stoichiometry Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry 11 Stoichiometry Test
 - Setting Reading Goals Chemistry 11 Stoichiometry Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry 11 Stoichiometry Test
 - Fact-Checking eBook Content of Chemistry 11 Stoichiometry Test
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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