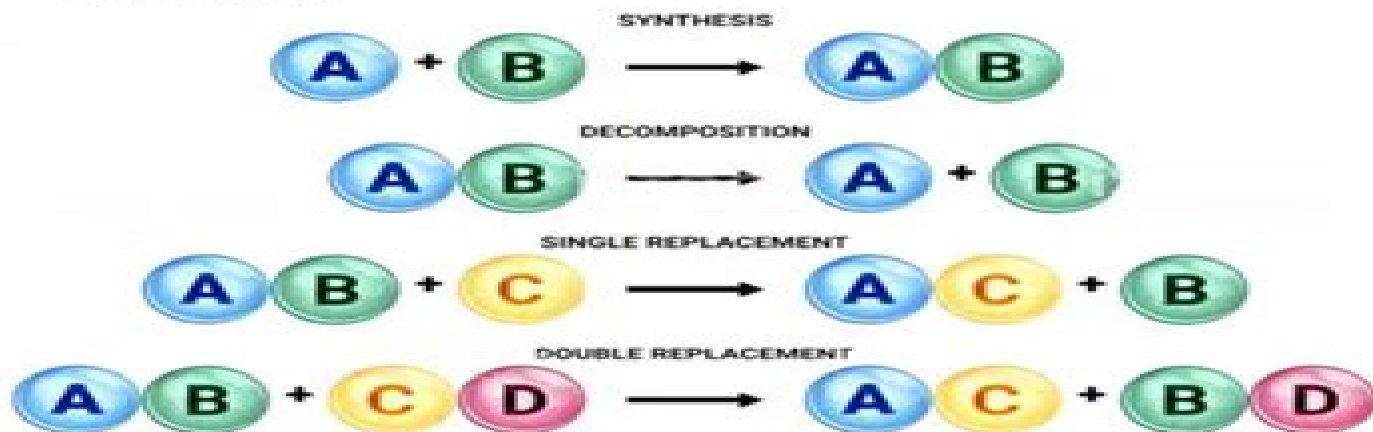


III. Types of Chemical Reactions

The 4 primary types of chemical reactions are:

1. **Combination (synthesis) reaction:** Two or more reactants combine to make a new product. Going from simple to complex molecules.
2. **Decomposition reaction:** A single reactant breaks down to form two or more products. Going from complex to simple molecules.
3. **Single Displacement reaction:** A single element replaces a similar element of an adjacent reactant compound.
4. **Double Displacement reaction:** Two compounds exchange elements, producing two new compounds.



ACTIVITY

Identify the type of reaction for the following equations.

Chemical Reaction	Type of reaction
$\text{Cu}_{(s)} + 2\text{AgNO}_{3(aq)} \rightarrow \text{Cu}(\text{NO}_3)_{2(aq)} + 2\text{Ag}_{(s)}$	
$2\text{KI}_{(aq)} + \text{Pb}(\text{NO}_3)_{2(aq)} \rightarrow 2\text{KNO}_{3(aq)} + \text{PbI}_{2(s)}$	
$2\text{H}_2\text{O}_{2(l)} \rightarrow 2\text{H}_2\text{O}_{(l)} + \text{O}_{2(g)}$	
$\text{C}_6\text{H}_{12}\text{O}_6 \rightarrow 2\text{C}_2\text{H}_5\text{OH} + 2\text{CO}_2$	
$\text{Mg}_{(s)} + \text{CuSO}_{4(aq)} \rightarrow \text{MgSO}_{4(aq)} + \text{Cu}_{(s)}$	
$2\text{C}_8\text{H}_{18} + 25\text{O}_2 \rightarrow 16\text{CO}_2 + 18\text{H}_2\text{O}$	

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Pamela Mills,Amina El-Ashmawy



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learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

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