

Chemistry Final Exam Review # 3

Chapter 12: Chemical Bonding

Define the following terms: *Bond, ionic bonding, ionic compound, covalent bonding, nonpolar covalent bond, polar covalent bond, electronegativity, dipole moment, Lewis structure, octet rule, bonding pair, lone pair (unshared pair), single bond, double bond, triple bond, Valence shell electron pair repulsion theory (VSEPR), tetrahedral structure, Pyramidal structure, bent structure, trigonal planar structure, linear structure*

Questions:

1. How are ionic bonds and covalent bonds different?

Ionic bonds involve a transfer of electrons between atoms, while covalent bonds involve a sharing of electrons between atoms. Ionic bonds usually occur between a metal and a nonmetal, while covalent bonds occur between nonmetals.

2. How is a polar covalent bond different from a nonpolar covalent bond?

In polar covalent bonds, the electrons are not shared equally. The bonding electrons are closer to the atom with the higher electronegativity. Nonpolar covalent bonds involve equal sharing of electrons

3. For each of the following sets of elements, place the elements in order of increasing electronegativity.

- | | |
|--------------|------------------|
| a. K, Sc, Ca | <i>K, Ca, Sc</i> |
| b. Br, F, At | <i>At, Br, F</i> |
| c. C, O, N | <i>C, N, O</i> |

4. What type of ion do oxygen atoms typically form and why? (*Hint: Consider oxygen's electron configuration.*)

Oxygen atoms typically form ions with a charge of -2 (O^{2-}). Oxygen gains two electrons to become an ion in order to fill its outer energy level with electrons. In its electron configuration, $1s^2 2s^2 2p^4$, two additional electrons will fill the 2^{nd} energy level.

5. Predict the formula for an ionic compound formed from each pair of elements.

- | | |
|-----------|-------------------------|
| a. Mg, S | <i>MgS</i> |
| b. K, Cl | <i>KCl</i> |
| c. Cs, F | <i>CsF</i> |
| d. Ba, Br | <i>BaBr₂</i> |

Chemical Bonding Final Review Guide

**John Krenos, Joseph Potenza, Loretta
Jones, Lynn Koplitz, Thomas Spence**



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