

Chemistry Formulas

Dilution

$$M_1V_1 = M_2V_2$$

Energy of a photon

$$E_{\text{photon}} = hf = hc/\lambda$$

Pressure

$$P = F/A$$

Average Kinetic Energy

$$KE_{\text{avg}} = \frac{3}{2}RT$$

Ideal Gas Law

$$PV = nRT$$

Combined Gas Law

$$\frac{P_1V_1}{n_1T_1} = \frac{P_2V_2}{n_2T_2}$$

Dalton's Law of Partial Pressures

$$P_{\text{total}} = P_A + P_B + \dots$$

Dalton's Law of Partial Pressures

$$P_A = \chi_A P_{\text{total}}$$

Graham's Law of Effusion

$$\frac{r_A}{r_B} = \sqrt{\frac{M_B}{M_A}}$$

Henry's Law

$$P_A = k_H[A]$$

Freezing Point Depression

$$\Delta T_f = -iK_f M$$

Boiling Point Elevation

$$\Delta T_b = iK_b m$$

Vapor Pressure Depression (Raoult's Law)

$$P_{\text{soln}} = \chi_{\text{solvent}} P_{\text{solvent}}^0$$

Osmotic Pressure (π)

$$\pi = iMRT$$

Arrhenius Equation

$$k = Ae^{-E_a/RT}$$

Equilibrium Constant Expressions

$$K_c = \frac{[\text{products}]}{[\text{reactants}]}$$

Equilibrium Constant Expressions

$$K_{eq} = \frac{k_{\text{forward}}}{k_{\text{reverse}}}$$

Equilibrium Constant Expressions

$$K_p = \frac{P_{\text{products}}}{P_{\text{reactants}}}$$

Reaction Quotient (Q)

$$Q = \frac{[\text{products}]}{[\text{reactants}]}$$

Solubility Product Constant (K_{sp})

$$K_{sp} = \frac{[\text{products}]}{[\text{reactants}]}$$

pH & pOH

$$pH = -\log[H^+]$$

pH & pOH

$$pOH = -\log[OH^-]$$

pH & pOH

$$pH + pOH = 14$$

pK_a & pK_b

pK_a & pK_b

pK_a & pK_b

Neutralization Reaction

Buffers

Enthalpy of Formation

First Law of Thermodynamics

Pressure Volume Work

Calorimetry Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Heat Curves & Thermal Energy (q)

Entropy (S)

Bond Dissociation Energy

Gibb's Free Energy (ΔG)

Gibb's Free Energy (ΔG)

Gibb's Free Energy (ΔG)

Standard Cell Potential

Standard Cell Potential

Nernst Equation

Kinetics (always 1st order)

Kinetics (always 1st order)

Kinetics (always 1st order)

Nuclear Binding Energy

$$pK_a = -\log[K_a]$$

$$pK_b = -\log[K_b]$$

$$pK_a + pK_b = 14$$

$$n_A M_A V_A = n_B M_B V_B$$

$$pH = pK_a + \log[A^-]/[HA]$$

$$\Delta H_f = \sum n \Delta H_f^{\circ}(\text{product}) - \sum n \Delta H_f^{\circ}(\text{reactants})$$

$$\Delta E = q + w$$

$$w = -P\Delta V$$

$$q = -C_{\text{calorimeter}}\Delta T$$

$$q = mc\Delta T$$

$$q = m\Delta H_{\text{fusion}}$$

$$q = m\Delta H_{\text{vaporization}}$$

$$\Delta S = \ln S_{\text{products}} - \ln S_{\text{reactants}}$$

$$\Delta H = \sum \Delta H_{\text{reactants}} - \sum \Delta H_{\text{products}}$$

$$\Delta G^{\circ} = \Delta H^{\circ} - T\Delta S^{\circ}$$

$$\Delta G = \Delta G^{\circ} + RT \ln Q$$

$$\Delta G^{\circ} = -RT^{\circ} \ln(K_{eq})$$

$$E^{\circ} = E^{\circ}_{\text{reduction}} + E^{\circ}_{\text{oxidation}}$$

$$E^{\circ} = E^{\circ}_{\text{cathode}} + E^{\circ}_{\text{anode}}$$

$$E_{\text{cell}} = E^{\circ} - \frac{0.0592}{n} \log Q$$

$$N = N_0 e^{-kt}$$

$$\ln N = \ln N_0 - kt$$

$$t_{1/2} = \frac{0.693}{k}$$

$$E = \Delta mc^2$$

Chemical Names And Formulas Chemistry Study Guide

RJ Shavelson



Chemical Names And Formulas Chemistry Study Guide:

This book delves into Chemical Names And Formulas Chemistry Study Guide. Chemical Names And Formulas Chemistry Study Guide is a vital topic that must be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Chemical Names And Formulas Chemistry Study Guide, encompassing both the fundamentals and more intricate discussions.

1. The book is structured into several chapters, namely:
 - Chapter 1: Introduction to Chemical Names And Formulas Chemistry Study Guide
 - Chapter 2: Essential Elements of Chemical Names And Formulas Chemistry Study Guide
 - Chapter 3: Chemical Names And Formulas Chemistry Study Guide in Everyday Life
 - Chapter 4: Chemical Names And Formulas Chemistry Study Guide in Specific Contexts
 - Chapter 5: Conclusion
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 3. In chapter 2, this book will delve into the foundational concepts of Chemical Names And Formulas Chemistry Study Guide. This chapter will elucidate the essential principles that must be understood to grasp Chemical Names And Formulas Chemistry Study Guide in its entirety.
 4. In chapter 3, this book will examine the practical applications of Chemical Names And Formulas Chemistry Study Guide in daily life. The third chapter will showcase real-world examples of how Chemical Names And Formulas Chemistry Study Guide can be effectively utilized in everyday scenarios.
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 6. In chapter 5, this book will draw a conclusion about Chemical Names And Formulas Chemistry Study Guide. The final chapter will summarize the key points that have been discussed throughout the book.
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