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$c = 3.00 \times 10^8 \text{ m/s}$ $h = 6.63 \times 10^{-34} \text{ J}\cdot\text{s}$ Avogadro's Number = 6.02×10^{23}

The ionization energy for hydrogen, if the electron is in the $n = 1$ main energy level, is $2.18 \times 10^{-18} \text{ J}$.

$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^{14} 6d^{10}$

1. An aqueous solution is prepared by dissolving 20.0 grams of potassium sulfate, K_2SO_4 (FW = 174) in sufficient water to prepare 500 milliliters of solution. What is $[\text{K}^+]$, the molarity of K^+ ion, in this solution?

1. 0.115 M
2. 0.230 M
3. 0.460 M
4. 0.080 M

$$[\text{K}_2\text{SO}_4] = \frac{(20.0)}{(174)} = 0.230 \text{ M}$$



$$[\text{K}^+] = 2(0.230) = \underline{\underline{0.460 \text{ M}}}$$

2. At 25.0 °C, the pOH of a solution is 4.75, what is $[\text{H}^+]$ in the solution?

1. $[\text{H}^+] = 9.25 \text{ M}$
2. $[\text{H}^+] = 1.78 \times 10^{-5} \text{ M}$
3. $[\text{H}^+] = 5.62 \times 10^{-10} \text{ M}$
4. $[\text{H}^+] = 9.61 \times 10^{-5} \text{ M}$

$$\text{pH} = 14 - 4.75 = 9.25$$

$$[\text{H}^+] = 10^{-9.25} = \underline{\underline{5.62 \times 10^{-10} \text{ M}}}$$

3. A compound of carbon and hydrogen is combusted completely with O_2 to produce CO_2 and H_2O . In a particular experiment, the combustion of a sample of the compound produces 5.50 grams of CO_2 and 4.50 grams of water. What is the mass of hydrogen in the sample that was combusted?

1. 1.50 grams H
2. 0.500 gram H
3. 0.250 gram H
4. 1.00 gram H

$$\frac{4.50 \text{ g H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol H}_2\text{O}} \times \frac{2 \text{ mol H}}{1 \text{ mol H}_2\text{O}} \times \frac{1.0 \text{ g H}}{1 \text{ mol H}} = \underline{\underline{0.500 \text{ g H}}}$$

Chem 1035 Placement Test

VM Jensen



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