Balancing Redox Reactions Worksheet

Acid Solutions

$$Mn^{2+} + BiO_3^- ===> MnO_4^- + Bi^{3+}$$

 $2 (4 H_2O + Mn^{2+} ===> MnO_4^- + 8 H^+ 5 e)$
 $5 (2 e + 6 H^+ + BiO_3^- ===> Bi^{3+} + 3 H_2O)$
 $14 H^+ + 2 Mn^{2+} + 5 BiO_3^- ===> 2 MnO_4^- + 5 Bi^{3+} + 7 H_2O$

$$ClO_3$$
 + Cl ===> Cl_2 + ClO_2
2 (e + 2 H + + ClO_3 ===> ClO_2 + H_2O)
2 Cl ===> Cl_2 + 2 e)
4 H + 2 ClO_3 + 2 Cl ===> 2 ClO_2 + 2 H_2O + Cl_2

$$P + Cu^{2+} ===> Cu + H_2PO_4^-$$

 $2 (4 H_2O + P ===> H_2PO_4^+ + 6 H^+ + 5 c)$
 $5 (2 e + Cu^{2+} ===> Cu)$
 $8 H_2O + 2 P + 5 Cu^{2+} ===> 2 H_2PO_4^+ + 12 H^+ + 5 Cu)$

$$PH_3 + I_2 ===> H_3PO_2^{-} + I^{-}$$

 $2 (2 H_2O + PH_3 ===> H_3PO_2^{-} + 4 H^{+} + 3 e)$
 $3 (2 e + I_2 ===> 2 I^{-})$
 $4 H_2O + 2 PH_3 + 3 I_2 ===> 2 H_3PO_2^{-} + 8 H^{+} + 6 I^{-}$

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Harold P. C. Burrell

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