

CHEM 101 Quiz #6 Chapter 6 ANSWERS**1. (4 points)**

The specific heat (s) of iron (Fe) is 0.40 J/(g K) . Calculate the heat effect q when 20 g of iron is heated from 100°C to 200°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \text{ in } ^\circ\text{C} \quad q = 20 \times 0.40 \times (200 - 100) = +800 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Mg = 24.0 g/mol .

Calculate the enthalpy change when 2.40 g Mg(s) reacts with excess oxygen. Include the sign!

$$2.40 \text{ g Mg} = 2.40/24.0 = 0.1 \text{ mol} \rightarrow \Delta H = 0.1(\text{mol}) \times (-600)(\text{kJ/mol}) = -60 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following four reactions is a *formation reaction*? Circle the correct answer.



Look up the definition of a formation reaction in Chang or in your lecture notes: the compound is formed from its elements, and the elements must be as they naturally occur so e.g. Cl_2 not Cl , and H_2 not H !

CHEM 101 Section L52 Quiz #6**1. (4 points)**

The specific heat (s) of copper (Cu) is 0.38 J/(g K) . Calculate the heat effect q when 10 g of copper is heated from 100°C to 300°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \text{ in } ^\circ\text{C} \quad q = 10 \times 0.38 \times (300 - 100) = +760 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Ca = 40.0 g/mol .

Calculate the enthalpy change ΔH when 8.0 g Ca(s) reacts with excess oxygen. Include the sign!!

$$8.0 \text{ g Ca} = 8.0/40 = 0.2 \text{ mol} \rightarrow \Delta H = 0.2(\text{mol}) \times (-650)(\text{kJ/mol}) = -130 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following reactions is a *formation reaction*? Circle the correct answer.



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Disha Experts

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The Hidden Curriculum - Faculty Made Tests in Science Sheila Tobias,1997

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Roger Slade, Maureen Slade,2011-04 Study as you go with Cambridge Checkpoints Updated annually to provide the most up to date exam preparation available Cambridge Checkpoints provides everything you need to prepare for your exams in a go anywhere format that fits easily into your school bag Recent official exam papers with suggested responses Hundreds of additional past exam and exam style questions with answers Dot point summaries of key topics and concepts to help you pinpoint where you need further revision

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Joseph A. Mascetta, Mark Kernion,2019-08-06 A self teaching guide for students Chemistry The Easy Way provides easy to

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