

Stoichiometry Lab - **ANSWERS****PRE-LAB QUESTIONS:**

1. Calculate the molar mass of sodium bicarbonate (NaHCO_3). 84.003 g/mol
 2. Calculate the molar mass of acetic acid (CH_3COOH). 60.052 g/mol
 3. Calculate the molar mass of carbon dioxide (CO_2). 44.003 g/mol
 4. Vinegar is generally 4-5% acetic acid dissolved in distilled water. That means that a 5% vinegar solution contains 0.05 mol of acetic acid per liter of solution. How many moles of acetic acid are present in 30.00 mL?
- $$\frac{\text{mols acetic acid}}{\text{L soln}} = \frac{\text{mols acetic acid}}{\text{L soln}} \quad \text{X} = \underline{0.0015} \text{ mols } \text{CH}_3\text{COOH} \text{ in } 30.0 \text{ mL of vinegar}$$
5. When baking soda (sodium bicarbonate) reacts with vinegar (acetic acid), they form sodium acetate, water, and carbon dioxide gas. Balance the chemical equation for this reaction.



6. A student measures out 5.00 g of baking soda to react with 50.00 mL of vinegar (which contains 0.042 moles of acetic acid).

a) Determine the limiting reagent, excess reagent, and theoretical yield (in grams) of carbon dioxide. Show your work.

$$5.00 \text{ g baking soda} \times \frac{1 \text{ mol baking soda}}{84.003 \text{ g baking soda}} = \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol baking soda}} \times \frac{44.003 \text{ g carbon dioxide}}{1 \text{ mol carbon dioxide}} = 2.62 \text{ g carbon dioxide}$$

$$0.042 \text{ mols acetic acid} \times \frac{1 \text{ mol carbon dioxide}}{1 \text{ mol acetic acid}} = \frac{44.003 \text{ g carbon dioxide}}{1 \text{ mol carbon dioxide}} = 1.86 \text{ g carbon dioxide}$$

*** Since acetic acid produced the least amount of product, the acetic acid in vinegar is the **limiting reagent** and the baking soda is the **excess reagent**.

*** The theoretical yield of carbon dioxide is 1.86 g.

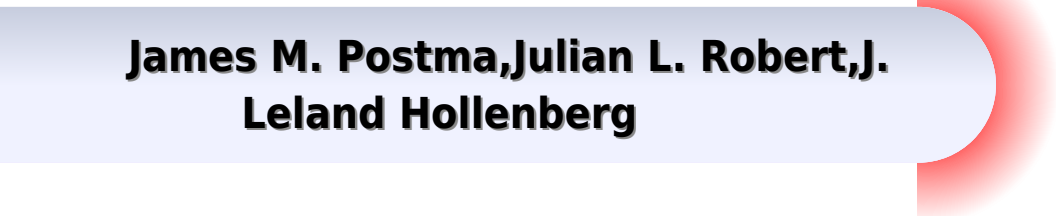
- b) This reaction will produce carbon dioxide bubbles that escape. After the reaction, the mass of the reaction decreased by 1.50 g. This is the actual mass of carbon dioxide that was produced from the reaction. Calculate the percentage yield of carbon dioxide. Show your work.

$$\text{Percent Yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$\text{Percent Yield} = \frac{1.50 \text{ g}}{1.86 \text{ g}} \times 100 = 79.6\% \text{ yield}$$

Chemistry Stoichiometry Lab Answers

**James M. Postma, Julian L. Robert, J.
Leland Hollenberg**



Chemistry Stoichiometry Lab Answers:

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Working with Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Julie Ellefson Kuehn, 2004-02-20 With this modular laboratory program students build skills using important chemical concepts and techniques to the point where they are able to design a solution to a scenario drawn from a professional environment The scenarios are drawn from the lives of people who work with chemistry every day ranging from field ecologists to chemical engineers and include many health professionals as well

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Sustainable Green Chemistry Mark Anthony Benvenuto, 2017-04-10 *Sustainable Green Chemistry* the 1st volume of Green Chemical Processing covers several key aspects of modern green processing The scope of this volume goes beyond bio and organic chemistry highlighting the ecological and economic benefits of enhanced sustainability in such diverse fields as petrochemistry metal production and wastewater treatment The authors discuss recent progresses and challenges in the implementation of green chemical processes as well as their transfer from academia to industry and teaching at all levels Selected successes in the greening of established processes and reactions are presented including the use of switchable polarity solvents actinide recovery using ionic liquids and the removal of the ubiquitous bisphenol A molecule from effluent streams by phytodegradation

Science Education in the 21st Century Tang Wee Teo, Aik-Ling Tan, Yann Shiou Ong, 2020-06-29 This book reflects on science education in the first 20 years of the 21st century in order to promote academic dialogue on science education from various standpoints and highlights emergent new issues such as education in science education research It also defines new research agendas that should be moved forward and inform new trajectories through the rest of the century Featuring 21 thematically grouped chapters it includes award winning papers and other significant papers that address the theme of the 2018 International Science Education Conference

Regents Exams and Answers: Chemistry--Physical Setting Revised Edition Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day

Integrating Green and Sustainable Chemistry Principles into Education Andrew P. Dicks, Loyd D. Bastin, 2019-07-19 *Integrating Green and Sustainable Chemistry Principles into Education* draws on the knowledge and experience of scientists and educators already working on how to encourage green chemistry integration in their teaching both within and outside of academia It highlights

current developments in the field and outlines real examples of green chemistry education in practice reviewing initiatives and approaches that have already proven effective By considering both current successes and existing barriers that must be overcome to ensure sustainability becomes part of the fabric of chemistry education the book s authors hope to drive collaboration between disciplines and help lay the foundations for a sustainable future Draws on the knowledge and expertise of scientists and educators already working to encourage green chemistry integration in their teaching both within and outside of academia Highlights current developments in the field and outlines real examples of green chemistry education in practice reviewing initiatives and approaches that have already proven effective Considers both current successes and existing barriers that must be overcome to ensure sustainability *Chemistry in the Laboratory* James M. Postma, Julian L. Robert, J. Leland Hollenberg, 2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations **Instructor's Guide for Introductory Chemistry in the Laboratory** James F. Hall, 1996 Nelson Chemistry Michael Dzwiniel, Frank Jenkins, Dick Tompkins, Hans Van Kessel, 1992 Basic resource for Senior High Chemistry 20 30 ca 1993 2007

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Table of Contents Chemistry Stoichiometry Lab Answers

1. Understanding the eBook Chemistry Stoichiometry Lab Answers
 - The Rise of Digital Reading Chemistry Stoichiometry Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Stoichiometry Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Stoichiometry Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Stoichiometry Lab Answers
 - Personalized Recommendations
 - Chemistry Stoichiometry Lab Answers User Reviews and Ratings
 - Chemistry Stoichiometry Lab Answers and Bestseller Lists

5. Accessing Chemistry Stoichiometry Lab Answers Free and Paid eBooks
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 - Chemistry Stoichiometry Lab Answers Budget-Friendly Options
6. Navigating Chemistry Stoichiometry Lab Answers eBook Formats
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 - Chemistry Stoichiometry Lab Answers Enhanced eBook Features
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 - Highlighting and Note-Taking Chemistry Stoichiometry Lab Answers
 - Interactive Elements Chemistry Stoichiometry Lab Answers
8. Staying Engaged with Chemistry Stoichiometry Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Stoichiometry Lab Answers
9. Balancing eBooks and Physical Books Chemistry Stoichiometry Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Stoichiometry Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Stoichiometry Lab Answers
 - Setting Reading Goals Chemistry Stoichiometry Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Stoichiometry Lab Answers
 - Fact-Checking eBook Content of Chemistry Stoichiometry Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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