

HSC Chemistry Notes:

- Module 5: Equilibrium and Acid Reactions
- Module 6: Acid and Base Reactions
- Module 7: Organic Chemistry
- Module 8: Applying Chemical Ideas

Module 5: Equilibrium and Acid Reactions

Static and Dynamic Equilibrium

Reversibility of chemical reactions:

Formation of hydrated and dehydrated cobalt (II) chloride:

Dehydrated cobalt (II) chloride + water $\rightleftharpoons$ hydrated cobalt (II) chloride



- When the hydrated cobalt (II) chloride is heated, it becomes blue again as the water evaporated.

Reaction of Iron (III) nitrate and potassium thiocyanate:

Iron (III) nitrate + potassium thiocyanate $\rightleftharpoons$ iron (III) thiocyanate



- When $(\text{Fe}(\text{NO}_3)_3)$ and (KSCN) are mixed, they form a red-coloured ion, iron (III) thiocyanate
- Is a reversible reaction and forms an equilibrium

Burning Magnesium:



- An irreversible combustion reaction, releasing energy into the system, producing white magnesium oxide.

Burning Steel Wool:



- An irreversible reaction and is a combustion reaction.
- As the steel wool caught on fire and went red with sparks flying out.

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