

THE MOLE AND AVOGADRO'S NUMBER

Name _____

One mole of a substance contains Avogadro's Number (6.02×10^{23}) of molecules. (atoms)

How many molecules are in the quantities below?

1. 2.0 moles $x = 2.0 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{1.2 \times 10^{24} \text{ molecules}}$

2. 1.5 moles $x = 1.5 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{23} \text{ molecules}}$

3. 0.75 mole $x = 0.75 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{4.5 \times 10^{23} \text{ molecules}}$

4. 15 moles $x = 15 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{9.0 \times 10^{24} \text{ molecules}}$

5. 0.35 mole $x = 0.35 \text{ mol} \times \left(\frac{6.02 \times 10^{23} \text{ molecules}}{1 \text{ mol}} \right) = \underline{2.1 \times 10^{23} \text{ molecules}}$

How many moles are in the number of molecules below?

1. 6.02×10^{23} $x \text{ moles} = 6.02 \times 10^{23} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.0 \text{ mol}}$

2. 1.204×10^{24} $x \text{ moles} = 1.204 \times 10^{24} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.0 \text{ mol}}$

3. 1.5×10^{20} $x \text{ moles} = 1.5 \times 10^{20} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{2.5 \times 10^{-4} \text{ mol}}$ (0.00025 mol)

4. 3.4×10^{26} $x \text{ moles} = 3.4 \times 10^{26} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{565 \text{ mol}}$

5. 7.5×10^{19} $x \text{ moles} = 7.5 \times 10^{19} \text{ molecules} \left(\frac{1 \text{ mol}}{6.02 \times 10^{23} \text{ molecules}} \right) = \underline{1.2 \times 10^{-4} \text{ mol}}$

Chemistry If8766 Pg 74 Answers

Yi-Tong Ma

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Table of Contents Chemistry If8766 Pg 74 Answers

1. Understanding the eBook Chemistry If8766 Pg 74 Answers
 - The Rise of Digital Reading Chemistry If8766 Pg 74 Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry If8766 Pg 74 Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry If8766 Pg 74 Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry If8766 Pg 74 Answers

- Personalized Recommendations
 - Chemistry If8766 Pg 74 Answers User Reviews and Ratings
 - Chemistry If8766 Pg 74 Answers and Bestseller Lists
5. Accessing Chemistry If8766 Pg 74 Answers Free and Paid eBooks
 - Chemistry If8766 Pg 74 Answers Public Domain eBooks
 - Chemistry If8766 Pg 74 Answers eBook Subscription Services
 - Chemistry If8766 Pg 74 Answers Budget-Friendly Options
 6. Navigating Chemistry If8766 Pg 74 Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry If8766 Pg 74 Answers Compatibility with Devices
 - Chemistry If8766 Pg 74 Answers Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry If8766 Pg 74 Answers
 - Highlighting and Note-Taking Chemistry If8766 Pg 74 Answers
 - Interactive Elements Chemistry If8766 Pg 74 Answers
 8. Staying Engaged with Chemistry If8766 Pg 74 Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry If8766 Pg 74 Answers
 9. Balancing eBooks and Physical Books Chemistry If8766 Pg 74 Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry If8766 Pg 74 Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry If8766 Pg 74 Answers
 - Setting Reading Goals Chemistry If8766 Pg 74 Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry If8766 Pg 74 Answers

- Fact-Checking eBook Content of Chemistry If8766 Pg 74 Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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