

Study Guide – Chapter 10: Gases

Test Date: _____

Chapter Ten:

- Kinetic Molecular Theory
 - 5 statements that describe the behavior of "ideal" gases
 - Ideal gases do not actually exist however real gases do behave almost ideally, so we assume the 5 statements (as well as the gas laws) to be true
 - real gases behave almost ideally at high temperatures and low pressures (this can be explained by the characteristics of gases as high temps and low pressures and relating it to the 5 statements)
 - The 5 statements all can relate back to the motion/characteristics of gases
 - moving extremely quickly in random motion
 - particles are extremely far away from one another, most of the volume that a gas takes up is empty space
- Pressure
 - conversions between Pa, kPa, torr, mm Hg, and atm
- The Gas Laws
 - STP – standard temperature and pressure values
 - Charles's Law: temperature & volume change
 - Boyle's Law: pressure & volume change
 - Gay-Lussac's Law: temperature & pressure
 - Avogadro's Law: volume & moles change
 - Avogadro's Hypothesis – leads to molar volume relationship
 - 1 mole of ANY gas = 22.4 L @ STP only
 - Combined Gas Law: P, V, & T changing conditions
 - Ideal Gas Law: $PV = nRT$
 - can be rearranged to solve for density, mass, or molar mass of a gas
- Gas Mixtures
 - all gases mix homogeneously (again it relates to the characteristics/motion of gases)
 - Dalton's Law of Partial Pressures
 - $P_t = P_1 + P_2 + P_3 \dots$ ect (also known the rearrangements of this formula, like the mole fraction formula or using total # of moles in the ideal gas formula to find P_t)
 - Collecting Gases over Water
 - this method for collecting a gas will have both the gas you are collecting as well as water vapor mixed together
 - $P_t = P_G + P_{H_2O}$
 - you must know the pressure of the water vapor so you can find the pressure of just the gas by itself (@ a specific temp that pressure will be given on a chart or in the problem)
- Effusion & Diffusion
 - know what each process is by definition
 - how does the KE of a gas relate to temperature? (also part of the 5 KMT statements)
 - KE is also related to mass and velocity, thus at a given temperature a gas with a smaller molar mass will move faster and vice versa

Chemistry Gases Study Guide

Thomas Griffiths



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