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CHEMISTRY: COUNTING ATOMS IN COMPOUNDS WORKSHEET #7.0.1

INSTRUCTIONS: Write the quantity of atoms of each element opposite the formula of the compound for the quantity of formula units and molecules shown:

For example: $5\text{P}_2\text{O}_5$ $\text{P} = (5 \times 2 =) 10$ $\text{O} = (5 \times 3 =) 15$

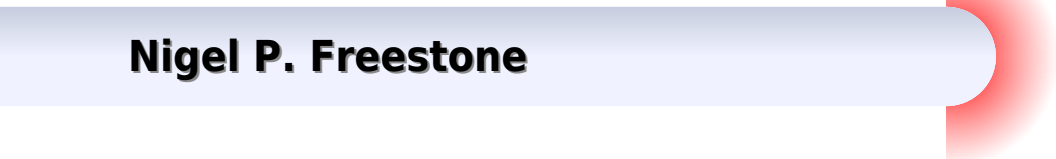
For example: $4\text{Zn}(\text{NO}_3)_2$ $\text{Zn} = (4 \times 1 =) 4$ $\text{N} = (4 \times 1 \times 2 =) 8$ $\text{O} = (4 \times 3 \times 2 =) 24$

- | | | | | |
|-----|---|--|--|---|
| 1. | $4\text{K}_2\text{CO}_3$ | $\text{K} = \underline{\hspace{1cm}}$ | $\text{C} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ |
| 2. | $2\text{Sr}_3(\text{PO}_4)_2$ | $\text{Sr} = \underline{\hspace{1cm}}$ | $\text{P} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ |
| 3. | $3\text{N}_2\text{O}_{10}$ | $\text{N} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ | |
| 4. | $2(\text{NH}_4)_3\text{N}$ | $\text{N} = \underline{\hspace{1cm}}$ | $\text{H} = \underline{\hspace{1cm}}$ | |
| 5. | $8\text{Cl}_2\text{O}$ | $\text{Cl} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ | |
| 6. | $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$ | $\text{Ca} = \underline{\hspace{1cm}}$ | $\text{C} = \underline{\hspace{1cm}}$ | $\text{H} = \underline{\hspace{1cm}}$ $\text{O} = \underline{\hspace{1cm}}$ |
| 7. | 12NaBr | $\text{Na} = \underline{\hspace{1cm}}$ | $\text{Br} = \underline{\hspace{1cm}}$ | |
| 8. | $4\text{Al}(\text{OH})_3$ | $\text{Al} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ | $\text{H} = \underline{\hspace{1cm}}$ |
| 9. | 3NaHCO_3 | $\text{Na} = \underline{\hspace{1cm}}$ | $\text{H} = \underline{\hspace{1cm}}$ | $\text{C} = \underline{\hspace{1cm}}$ $\text{O} = \underline{\hspace{1cm}}$ |
| 10. | $5\text{Ga}_2(\text{Cr}_2\text{O}_7)_3$ | $\text{Ga} = \underline{\hspace{1cm}}$ | $\text{Cr} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ |
| 11. | $7\text{C}_2\text{S}_2$ | $\text{C} = \underline{\hspace{1cm}}$ | $\text{S} = \underline{\hspace{1cm}}$ | |
| 12. | $4\text{Fe}_2\text{O}_3$ | $\text{Fe} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ | |
| 13. | $6\text{Ba}(\text{MnO}_4)_2$ | $\text{Ba} = \underline{\hspace{1cm}}$ | $\text{Mn} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ |
| 14. | $3\text{V}_2\text{O}_5$ | $\text{V} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ | |
| 15. | 2KNO_3 | $\text{K} = \underline{\hspace{1cm}}$ | $\text{N} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ |
| 16. | 9MgSO_4 | $\text{Mg} = \underline{\hspace{1cm}}$ | $\text{S} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ |
| 17. | $5\text{Al}_2(\text{SiO}_3)_2$ | $\text{Al} = \underline{\hspace{1cm}}$ | $\text{Si} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ |
| 18. | $4\text{Au}(\text{IO}_3)_3$ | $\text{Au} = \underline{\hspace{1cm}}$ | $\text{I} = \underline{\hspace{1cm}}$ | $\text{O} = \underline{\hspace{1cm}}$ |

(Continued)

Chemistry Counting Atoms In Compounds

Nigel P. Freestone



Chemistry Counting Atoms In Compounds:

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percentage composition empirical and molecular formula The book also covers concentration its units volumetric analysis and the relationship between volume mass and moles of gases Counting Moles culminates in you taking a Mole Driving Test On passing this test you are issued with a Counting Moles Driving License that will give you all the confidence required to correctly answer all mole calculations *Basic Concepts of Chemistry* Leo J. Malone, Theodore O. Dolter, 2011-12-27 The 9th edition of Malone's *Basic Concepts of Chemistry* provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter WileyPLUS sold separately from text *The Practice of Chemistry* Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book's unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB *Introductory Chemistry* Darrell D. Ebbing, Albert E. Richardson, 1995

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Chemistry Counting Atoms In Compounds Book Review: Unveiling the Magic of Language

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