

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

## CHAPTER

## 15

## STUDY GUIDE FOR CONTENT MASTERY

## Solutions

### Section 15.1 What are solutions?

In your notebook, read about the characteristics of solutions.

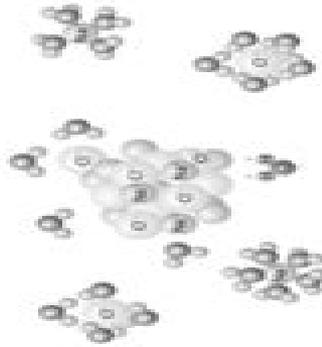
Use each of the terms below just once to complete the passage.

immiscible	liquid	soluble	solution
insoluble	miscible	solvent	solvent

Air is a(n) **(1)** solution of oxygen gas dissolved in nitrogen gas. The oxygen is an **(2)** solute and nitrogen is the **(3)** solvent. Because oxygen gas dissolves in a solvent, oxygen gas is an **(4)** soluble substance. A substance that does not dissolve is **(5)** insoluble. **(6)** Liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are **(7)** miscible. However, if one liquid is insoluble in another, the liquids are **(8)** immiscible.

Read about solvation in aqueous solutions in your notebook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



9. Hydration is solvation in which the solvent is water. What is solvation?  
**Solvation is the process of surrounding solute particles with solvent particles to form a solution.**

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## CHAPTER

## 15

## STUDY GUIDE FOR CONTENT MASTERY

## Section 15.1 continued

16. An sodium chloride dissolves in water. What happens to the sodium and chloride ions? **The sodium and chloride ions are separated and surrounded by the water molecules.**

water molecules.

17. Explain the orientation of the water molecules around the sodium ions and chloride ions. **Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.**

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? **How do you know?**

**The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.**

19. List three ways that the rate of solvation may be increased.

**stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent**

In your notebook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write *true* or *false*.

- true** 14. The overall energy change that occurs when a solution forms is called the heat of solution.

- false** 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

- false** 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

- true** 17. In a saturated solution, solvation and crystallization are in equilibrium.

- true** 18. Additional water can be dissolved in an unsaturated solution.

- true** 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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# Chemistry Content Mastery Study Guide Teacher Edition

**Jicheng Xie**



**Chemistry Content Mastery Study Guide Teacher Edition:**

## **Chemistry Content Mastery Study Guide Teacher Edition** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the energy of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such is the essence of the book **Chemistry Content Mastery Study Guide Teacher Edition**, a literary masterpiece that delves deep into the significance of words and their impact on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we will explore the book's key themes, examine its writing style, and analyze its overall effect on readers.

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