

APHY101- Anatomy & Physiology I

Chapter #2- Chemistry Comes Alive

Study Guide

- Define matter. What is the smallest unit of matter?

Matter = Anything that occupy space and has mass. Can exist in different forms: Solid, Liquid, or Gas.

- What is an element?

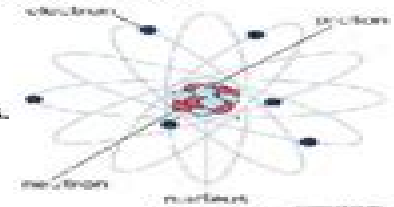
Element = a pure substance that cannot be created or broken down into simpler substances by ordinary chemical means. There're 118 recognized elements and 92 occur in nature, ex. Carbon, Oxygen, Hydrogen, Nitrogen (found most in human body). Every element is made up of building blocks called **ATOMS**.

- Define atomic number, atomic mass & atomic weight.

Atomic Number = the number of protons in the nucleus of an atom.

Atomic Mass = the sum of the number of neutrons and protons in an atomic nucleus.

Atomic Weight = The average mass of all the isotopes of an element.

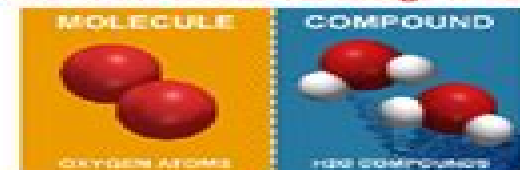


- Define atom

Atom = Smallest units of matter, no life properties (non-living), made up for every tiny Subatomic Particles.

- What are the three (3) subatomic particles that make up an atom & what is the location & charge of each?

- Nucleus** – Central core of an atom.
 - Protons (+): charged particles
 - Neutrons: Neutral; not charged
- Electrons** (-): charged particles that orbit the nucleus



- Define molecule

Molecule = A group of Atoms bonded together by chemical bonds. Can contain only one element or more.

- Define compound

Compound = Substance composed of two or more elements joined by chemical bonds.

For example, water (H₂O) has three atoms, two hydrogen (H) atoms and one oxygen (O) atom.

- All compounds are molecules, but not all molecules are compounds*



- Know the difference between an organic molecule/compound & an inorganic molecule/compound

Organic compounds always contain carbon while most inorganic compounds do not contain carbon.

- Why is water an important inorganic compound in the body?

Water is the most abundant inorganic compound, making up over 60% of the volume of cells and over 90% of body fluids like blood. Many substances dissolve in water and all the chemical reactions that take place in the body do so when dissolved in water. Other inorganic molecules help keep the acid/base balance (pH) and concentration of the blood and other body fluids stable.

Chemistry Comes Alive Study Guide

Jean Dorricott



Chemistry Comes Alive Study Guide:

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Chemistry John S. Phillips, Cheryl Wistrom, 2000 SCM Studyguide to Science and Religion Jean Dorricott, 2005 This undergraduate level one textbook provides an introduction to the apparently incompatible subjects of religion and science Each chapter contains references for finding out more about particular arguments be they scientific or religious areas for discussion Where particularly difficult concepts are referred to in the body of the text further explanations are provided in boxed sections

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