

Chapter 3 Chemical Reactions Practice Test

Name

Question 1. This is a list of words and terms used in your study of this topic. Match the term with the definition or description in the table below. The words and terms may be used once, more than once, or not at all.

Balanced equation	Exothermic reaction	Reactants
Coefficients	Hydrocarbons	Salt
Combustion	Law of conservation of mass	Word equation
Corrosion	Neutralization	
Endothermic reaction	Products	

Definition or description	Word or term (from the list above) which best fits the definition.
Atoms are not created or destroyed in a chemical reaction, only rearranged.	
A reaction with oxygen that releases heat and light, commonly by burning fuels.	
A chemical reaction that absorbs heat from the surroundings.	
The substances that are formed as the result of a chemical reaction.	
In a chemical reaction matter is neither created nor destroyed.	
Oxygen + hydrogen $\rightarrow$ water is an example of a:	
The substances that are written on the left hand side of the arrow in a chemical equation.	
A family of chemical substances that contain only carbon and hydrogen.	

(8 marks)

Question 2. The following statements are either true or false. Complete the table by writing T next to the true statements and F next to the false statements.

Statement	True or false?
The symbol H_2O means two atoms of hydrogen chemically bonded with one atom of oxygen to make one molecule of water.	
The gas produced when a metal reacts with an acid is hydrogen gas.	
Carbon dioxide is produced when a hydrocarbon such as methane, CH_4 , is burned.	
When ten molecules of methane are burned in oxygen, ten molecules of carbon dioxide would be formed.	
The products of combustion are a salt and water.	
Carbon dioxide gas can be produced from both reacting an acid with a carbonate and also by combusting propane, C_3H_8 .	

Chemistry Chemical Reactions Practice Test Answers

Laura Schellhardt



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