

## Chemistry 11

### Final Examination Review - Answers

#### Part A - True or False.

Indicate whether each of the following statements is true or false. Correct the false statements.

- F 1. The mass of an electron is equal to the mass of a proton.  
**The mass of an electron is less than the mass of a proton.**
- T 2. The mass of a proton is approximately equal to the mass of a neutron.
- T 3. The atomic number represents the number of protons in a nucleus.
- T 4. The proton has a mass of approximately 1 u.
- F 5. The difference in mass of isotopes of the same element is due to the different number of protons in the nucleus.  
**The difference in mass of isotopes of the same element is due to the different number of neutrons in the nucleus.**
- T 6. The isotope carbon-12 is used as the relative mass standard for the atomic mass scale.
- F 7. The mass of the most-common isotope of each element is listed on the periodic table.  
**The average mass of the naturally occurring isotopes of each element is listed on the periodic table.**

#### Part B - Multiple Choice

- D 1. A(n) \_\_\_ is used to represent a compound.  
 a) symbol                      b) equation                      c) subscript                      d) formula
- A 2. \_\_\_ atoms or groups of atoms are called ions.  
 a) charged                      b) diatomic                      c) neutral                      d) monatomic
- A 3. For the formula of a compound to be correct, the algebraic addition of the charges on the atoms or ions in the compound must add up to \_\_\_\_.  
 a) zero                      b) one                      c) two                      d) four
- D 4. Potassium bromide is an example of a(n) \_\_\_ compound.  
 a) molecular                      b) organic                      c) polyatomic                      d) ionic
- B 5. The only common polyatomic ion that has a positive charge is the \_\_\_ ion.  
 a) phosphate                      b) ammonium                      c) sulfate                      d) nitrate
- A 6. In the formula  $\text{H}_2\text{SO}_4$ , the number 4 would be called a(n) \_\_\_\_.  
 a) subscript                      b) oxidation number                      c) coefficient                      d) charge
- B 7. Which subatomic particle contributes the least to the mass of an atom?  
 a) nucleon                      b) electron                      c) proton                      d) neutron
- D 8. In the free, or uncombined, state the number of protons in the nucleus of an element must equal the \_\_\_\_.  
 a) mass number                      c) mass number - atomic number  
 b) number of neutrons in the nucleus                      d) number of electrons present

# Chemistry 11 Final Exam Review

**A. A. Frempong**



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