

Stoichiometry Worksheet and Key



1. How many moles of O_2 will be formed from 1.65 moles of KClO_3 ?

1.65 mol KClO_3	_____ mol O_2	= _____ mol O_2
	_____ mol KClO_3	

2. How many moles of KClO_3 are needed to make 3.50 moles of KCl ?

3.50 mol KCl		= _____ mol KClO_3

3. How many moles of KCl will be formed from 2.73 moles of KClO_3 ?

_____		= _____



4. How many moles of Fe_2O_3 are produced when 0.275 moles of Fe is reacted?

0.275 mol Fe		= _____ mol Fe_2O_3

5. How many moles of Fe_2O_3 are produced when 31.0 moles of O_2 is reacted?

_____		= _____

6. How many moles of O_2 are needed to react with 8.9 moles of Fe ?

_____		= _____



7. How many moles of O_2 are produced when 1.26 moles of H_2O is reacted?

8. How many moles of H_2O are needed to produce 55.7 moles of H_2 ?

9. If enough H_2O is reacted to produce 3.40 moles of H_2 , then how many moles of O_2 must have been made?
(a bit challenging, but just think about it and you can probably figure it out)

Chemistry Stoichiometry Problem Sheet 2 Key

W. Wayne Meinke, John Keenan Taylor



Chemistry Stoichiometry Problem Sheet 2 Key:

The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can't do chemistry if they can't do the math. The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving. The book's unique focus on actual chemical practice, extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors. This special PowerPoint tour of the text was created by Don Wink http://www.bfwpub.com/pdfs/wink/POCPowerPoint_Final.ppt 832KB

Basics for Chemistry David A. Ucko, 2013-09-24 Basics of Chemistry provides the tools needed in the study of General Chemistry such as problem solving skills, calculation methods and the language and basic concepts of chemistry. The book is designed to meet the specific needs of underprepared students. Concepts are presented only as they are needed and developed from the simple to the complex. The text is divided into 18 chapters each covering some particular aspect of chemistry such as matter, energy and measurement, the properties of atoms, description of chemical bonding, study of chemical change and nuclear and organic chemistry. Undergraduate students will find the book as a very valuable academic material.

Resources in Education, 1982-04

Essays in Contemporary Chemistry Gerhard Quinkert, 2001 In this volume internationally renowned chemists recount their roles in the progress of chemistry research toward elucidation of biological processes. Beginning with a historical perspective on the development of X-ray crystallography, the reader is regaled with first-hand accounts of research milestones. Included are descriptions of the cutting-edge nuclear magnetic resonance and electron paramagnetic resonance spectroscopic techniques, the dynamics of ultra-fast reactions and the central problem of molecular oxygen activation in biological processes. The roles of chiral auxiliaries in organic synthesis and of photochemistry in natural product synthesis are told and innovations in combinatorial synthesis approaches are described. Contemporary issues in molecular recognition and modulation of molecular function are addressed, concluding with a missive regarding how the frontiers of medical ethics may be breached by molecular manipulations. The contributors, who number among the finest scientists in the world including two Nobel Prize winners, are Peter B. Dervan, Jack D. Dunitz, Christian Griesinger, Jean-Marie Lehn, Thomas F. Prisner, Gerhard Quinkert, Peter G. Schultz, Helmut Schwarz, Dieter Seebach, and Ahmed Zewail. Additionally, there is a prologue by Albert Eschenmoser for whom this collection was conceived and an epilogue that contains facsimiles of notes from his landmark lecture, "Synthesis of Coenzyme B₁₂: A Vehicle for Teaching Organic Synthesis." This book is definitely a must for all who want to read or to read again where we stand in our chemical comprehension of the fascinating relationship between chemical structure and biological processes, how we got here and what the future might hold.

Analytical Chemistry: Key to Progress on National Problems W. Wayne Meinke, John Keenan Taylor, 1972

Chemistry, Student Study Guide John A. Olmsted, Gregory M. Williams, 2005-02-02

100% Pure Chemical Understanding Every morning many of us are energized by a cup of coffee Imagine if you were as energized by understanding the chemistry in your morning cup from the coffee trees which fill red coffee berries with caffeine and a variety of other chemical substances to the feathery crystals formed by the caffeine molecules to the decaffeinating machines which use liquid solvents to remove this stimulant from some of the beans Now that's real chemical understanding Olmsted and Williams Fourth Edition of Chemistry focuses on helping you see and think about the world and even your coffee as a chemist This text helps you understand how chemical phenomena are governed by what happens at the molecular level apply critical thinking skills to chemical concepts and problems and master the basic mathematical techniques needed for quantitative reasoning You'll see the world as chemists do and learn to appreciate the chemical processes all around us A Fourth Edition with a lot of new perks Revisions include a new early energy chapter revised coverage of bonding expanded coverage of intermolecular forces and increased coverage of multiple equilibria including polyprotic acids New pedagogy strengthens students critical thinking and problem solving skills Visual Summaries at the end of each chapter use molecular and diagrammatic visual elements to summarize essential skills concepts equations and terms eGrade Plus provides an integrated suite of teaching and learning resources including a complete online version of the text links between problems and relevant sections in the online text practice quizzes the Visual Tutor Interactive LearningWare problems and lab demos as well as homework management and presentation features for instructors

A Problem-Solving Approach to Aquatic Chemistry James N. Jensen, 2023-01-05 A Problem Solving Approach to Aquatic Chemistry Enables civil and environmental engineers to understand the theory and application of aquatic equilibrium chemistry The second edition of A Problem Solving Approach to Aquatic Chemistry provides a detailed introduction to aquatic equilibrium chemistry calculation methods for systems at equilibrium applications of aquatic chemistry and chemical kinetics The text directly addresses two required ABET program outcomes in environmental engineering chemistry including stoichiometry equilibrium and kinetics and material and energy balances fate and transport of substances in and between air water and soil phases The book is very student centered with each chapter beginning with an introduction and ending with a summary that reviews the chapter's main points To aid in reader comprehension important terms are defined in context and key ideas are summarized Many thought provoking discussion questions worked examples and end of chapter problems are also included Each part of the text begins with a case study a portion of which is addressed in each subsequent chapter illustrating the principles of that chapter In addition each chapter has an Historical Note exploring connections with the people and cultures connected to topics in the text A Problem Solving Approach to Aquatic Chemistry includes Fundamental concepts such as concentration units thermodynamic basis of equilibrium and manipulating equilibria Solutions of chemical equilibrium problems including setting up the problems and algebraic graphical and computer solution techniques Acid base equilibria including the concepts of acids and bases titrations and alkalinity and acidity Complexation including metals ligands

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