

THE FOLLOWING QUESTIONS ARE PRACTICE QUESTIONS FOR THE CSN CHEMISTRY PLACEMENT EXAM

The chemistry placement test is used to assess your preparation for General Chemistry I, CHEM 121 at CSN. The test is a standardized test that will examine your knowledge of chemistry and basic mathematical skills. You will be provided scratch paper, a Periodic Table of the Elements and a nonprogrammable scientific calculator. The following sample questions are intended only as a guide to show you the relative level of difficulty and example topics covered in the Chemistry Placement Exam. The actual exam will have a time limit of 45 minutes to answer 44 multiple choice questions. Click the following link to view a [Periodic Table](#).

MULTIPLE CHOICE. Choose the single choice that best answers the question.

- 1) How many total atoms are in the formula



A) 12 B) 14 C) 9 D) 8

- 2) How many electrons are in Br^- ?

A) 4 B) 7 C) 36 D) 34

- 3) What are the coefficients for the following reaction when it is properly balanced?



A) 2, 3, 6, 1 B) 6, 1, 3, 2
C) 2, 3, 1, 6 D) 2, 1, 1, 3

- 4) Given that $4 \text{ NH}_3 + 5 \text{ O}_2 \rightarrow 4 \text{ NO} + 6 \text{ H}_2\text{O}$, when 4.50 mol of H_2O are formed, the amount of NO formed is

A) 3.00 mol B) 6.75 mol
C) 1.50 mol D) 4.50 mol

- 5) What type of a reaction occurs when a sodium hydroxide solution is mixed with an acetic acid solution?

A) gas evolution
B) oxidation-reduction
C) acid-base neutralization
D) precipitation

- 6) Which of the following atoms has the smallest radius?

A) Ne B) O
C) Li D) B

- 7) What is the molecular geometry if you have a double bond, a single bond and 1 lone pair around the central atom?

A) trigonal pyramidal
B) linear
C) tetrahedral
D) bent

- 8) What volume of 12.0 M HCl is required to make 75.0 mL of 3.50 M HCl?

A) 0.560 mL B) 257 mL
C) 560. mL D) 21.9 mL

$$9) \frac{(2.80 \times 10^{-3})(1.28 \times 10^1)}{(3.20 \times 10^{-4})} = \underline{\hspace{2cm}}$$

A) 1.27×10^1 B) 2.0×10^2
C) 1.2×10^6 D) 1.12×10^2

- 10) Of the molecules below, only _____ is polar.

A) CH_4 B) SbF_5
C) CO_2 D) AsH_3

- 11) Which one of the following is a metalloid?

A) S B) Ge C) C D) Pb

- 12) A certain atom has a charge of -2 and has 16 protons and 16 neutrons in the nucleus. How many electrons does this atom possess?

A) 34 B) 14 C) 32 D) 18

- 13) How many moles of pyridine ($\text{C}_5\text{H}_5\text{N}$) are contained in 3.13 g of pyridine?

A) 0.00404 B) 25.3
C) 0.0396 D) 0.319

- 14) Solve the following equation for x:

$$4x - b = 2x + c$$

A) $c - b - 6$ B) $\frac{c + b}{2}$
C) $\frac{c + b}{6}$ D) $\frac{c - b}{2}$

Chemistry Placement Test Study Guide Texas Tech

KJ Lindholm-Leary



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General Catalog - Texas Tech University Texas Tech University, 1929

Resources in Education , 1999

Friedland/Relyea Environmental Science for AP* Andrew Friedland, Rick Relyea, David Courard-Hauri, 2011-02-15

Friedland Relyea Environmental Science for AP was specifically developed to meet the requirements of the AP Environmental Science course and the needs of its students and teachers This highly anticipated new textbook explores the science behind environmental science and involves students with the fundamental concepts and findings that inform environmental decision making at all levels from personal choices to national and international policy This site will be the source for periodic updates on this exciting project as it draws closer to publication For the latest developments or if you would like to be a part of this project as a reviewer or class tester please contact Carlise Stemberge

Inventary of Federal Energy-related Environment and Safety Research for FY 1979 , 1980

Energy Research Abstracts , 1988

Inventary of Federal Energy-related

Environment and Safety Research for ... , 1980

Solar Energy Update , 1984-05

Understanding Risk to Wildlife

from Exposures to Per- and Polyfluorinated Alkyl Substances (PFAS) Mark S. Johnson, Michael J. Quinn Jr., Marc A.

Williams, Allison M. Narizzano, 2021-04-22

Understanding Risk to Wildlife from Exposures to Per and Polyfluorinated Alkyl Substances PFAS provides the most recent summary of toxicity data relevant to mammals birds reptiles and amphibians and provides values for use in risk assessment applications Predicting the bioaccumulation of PFAS in terrestrial wildlife including humans has proven to be extremely complex As a group PFAS act differently than traditional non ionic organic molecules where PFAS can break down and reform whereas some are demonstrated to be extremely persistent Where sufficient data are provided this book establishes toxicity reference values TRVs which are derived to assist in characterizing environmental sources of contamination and making risk based decisions Features Provides toxicity reference values TRVs for vertebrates mammals birds amphibians for PFAS where sufficient data are available and includes objective supporting background information Assigns a level of confidence to each TRV to provide the risk assessor with an understanding of the relative uncertainty associated with each value Presents toxicity data in the formats of scatter diagrams and tables for quick review and assessment Provides TRVs relevant for screening and decision making This book serves as a useful aid for risk assessors and managers in those industries that have sites contaminated with PFAS consultants tasked with evaluating risks at such sites and staff at regulatory agencies at various governmental levels who need to know how much contamination is considered safe for wildlife It will also appeal to researchers with an interest in filling the gaps in the current toxicological data for PFAS exposure

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