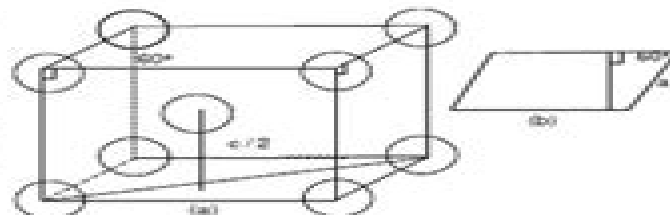


(6×4=24 Marks)

- Which of the following statement is(are) correct for the diamond structure?
 (A) Each atom has 4 nearest neighbors and 12 next nearest neighbour
 (B) The maximum proportion of the unit cell which may be filled by hard spheres is only 0.34.
 (C) The maximum proportion of the unit cell which may be filled by hard spheres is only 0.46.
 (D) None of these
- The mineral hawleyite. One form of CdS , crystallizes in one of the cubic lattices with edge length $5.87\text{\AA}$. The density of hawleyite is 4.63 g cm^{-3} .
 (i) In which lattice does hawleyite crystallize?
 (ii) Find the schottky defect in gm cm^{-3}
- If all the atoms touching one face plane are removed in solid A^+B^- having rock salt type structure, then calculate the formula of the compound and the defect brought by this removal.
- 100 ml 0.2 M NiCl_2 is treated with 100 ml 0.3 M dimethyl glyoxime then the mass of ppt (gm) formed will be ($\text{Ni}=58.7$)
- A yellow solid (A) is insoluble in water. It dissolves slowly in hot concentrated HNO_3 and a brown gas is evolved. Solid (A) gets dissolved in boiling Na_2SO_3 solution to form a clear solution (B). On acidification of this solution of (B), a colourless gas (C) with burning sulphur smell comes out along with white turbidity due to the formation of D. Which of the following statement/s is/are correct?
 (A) A and D may be similar compound
 (B) B is used in iodometric titration
 (C) B is used in removal of hardness of water
 (D) D solution in lyophobic colloid
- (i) A white precipitate (B) is formed when a mineral (A) is boiled with Na_2CO_3 solution
 (ii) the precipitate is filtered and the filtrate contains two compounds (C) and (D). The compound (C) is removed by crystallization and when CO_2 is passed through the mother liquor left (D) changes to (C).
 (iii) The compound (C) on strong heating gives two compounds (D) and (E).
 (iv) (E) on heating with cobalt oxide produces blue coloured substance (F)
 Identify (A) to (F) and give chemical equations for the reactions at steps (i) to (iv)

(2×8=16 Marks)

- Iron crystallizes in two bcc lattices the α -form below 910°C and the β form above 1400°C and in a fcc form between these two temperature of these three crystalline phases. Only the γ form can dissolve appreciable amount of carbon.
 a) What is the symmetry of a hole centered in a face of a body centered cubic unit cell and what is the max ratio of a foreign particle radius occupying such a hole to the host ion radius.
 b) What is the position of a hole whose coordinates are $(0, a/2, a/4)$ in a bcc unit cell and what is the max ratio of a foreign particle radius occupying such a hole to the host ion radius.
 c) Why does only γ form can dissolve appreciable amounts of carbon?
 - Potassium selenate is isomorphous with potassium sulphate and contains 50.0% of Se. What is the atomic weight of Se?
- The hexagonal close-packed lattice can be represented by figure. If $c = a\sqrt{8/3} = 1.633a$. There is an atom at each corner of the unit cell and another atom which can be located by moving one-third the distance along the diagonal of the rhombus base, starting at the lower left-hand corner and moving perpendicularly upward by $c/2$ figure. Mg crystallizes in this lattice and has a density of 1.74 g/cm^3 .
 (a) What is the volume of the unit cell?
 (b) What is a ?
 (c) What is the distance between nearest neighbors?
 (d) How nearest neighbor does each atom have?



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