

Equilibriumary Problems (Equilibrium):

1. balancing this equation
2. convert the moles of _____ at _____ grams _____ substance is
3. MAJOR TO MINOR substance _____ substance _____ that make partial
PERCENT BALANCED REACTIONS
4. convert the moles _____

Equilibrium: Best Case

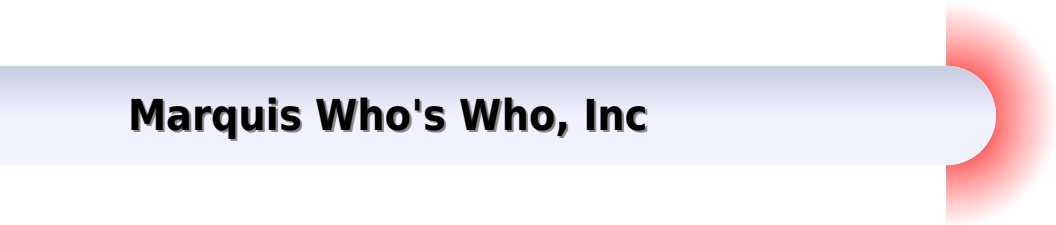
100.0 g C_2H_6 is (1 mol C_2H_6)(30.07 g C_2H_6)/mol = 30.07 mol C_2H_6 (1 mol C_2H_6 is (30.07 g C_2H_6)/30.07 mol C_2H_6) = [1 L/ 3.99 g] = 250.24 g C_2H_6

Equilibrium: Best Case: 100.0 g and substance is a liquid

100.0 g C_2H_6 is (1 mol C_2H_6)(30.07 g C_2H_6)/mol = 30.07 mol C_2H_6 (1 mol C_2H_6 is (30.07 g C_2H_6)/30.07 mol C_2H_6) = 30.07 mol C_2H_6 = 30.07 mol C_2H_6 = 30.07 mol C_2H_6

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Chemistry Note Taking Guide Episode 803 Book Review: Unveiling the Power of Words

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