



Chemistry

Paper Review and Study Guide

Name _____

Ionic & Metallic Compounds

1. Explain how the formation of ions contributes to the stability of ionic compounds.

2. Describe the characteristics of ionic compounds, including the electrostatic attraction between cations and anions.

3. Explain the difference between ionic and metallic bonds.

Name _____

Ionic Compounds

21. How does the charge of ions affect the strength of ionic bonds?

22. Give an example of an ionic compound.

23. What are interstitial alloys?

24. Name a property unique to ionic compounds.

25. What type of bond is formed in ionic compounds?

26. What is the relationship between ionic bond strength and the size of the ions?

27. How do you write the name for an ionic compound?

28. How do you write the formula for an ionic compound?

29. How do you write the name for a molecular compound?

30. How do you write the formula for a molecular compound?

Name _____

Ionic Compounds & Metals

Study Guide:

1. Chemical Bonds

- Definition:** The force that holds two atoms together.
- Types:** ionic bonds, covalent bonds, and metallic bonds.

The formation of ions is a fundamental concept in chemistry, playing a crucial role in the creation of ionic compounds. Atoms achieve stability by attaining a full outer shell of electrons, resembling the electron configuration of noble gases. This process involves gaining or losing electrons, leading to the formation of ions. Metals, with their tendency to lose electrons, become positively charged ions called cations. Conversely, nonmetals gain electrons to become negatively charged ions known as anions.

2. Ion Formation

- Cation:** An atom or molecule with a net electric charge due to the loss of one or more electrons.
- Anion:** A negatively charged ion (gain of electrons).

Sodium (Na), a metal, readily loses its single valence electron to achieve the stable electron configuration of neon (Ne), resulting in a sodium cation (Na⁺). This cation now possesses 11 protons and 10 electrons, leading to a net positive charge. The formation of an ion is driven by the desire to achieve a stable electronic configuration, mimicking the noble gases, which are known for their chemical inertness.

3. Ionic Compounds

Definition: Compounds formed from the electrostatic attraction between oppositely charged ions.

Formation: The transfer of electrons from a metal atom to a non-metal atom results in the formation of a cation and an anion. These oppositely charged ions are then attracted to each other, leading to the creation of an ionic bond. The resulting ionic compounds are characterized by high melting points and solubility in water.

4. Metallic Bonds

Definition: Bonds formed between metal atoms, where the valence electrons are delocalized and act as a 'sea of electrons' that holds the metal cations together.

Formation: Occurs between metals (which lose electrons) and non-metals (which gain electrons). The resulting structure is a lattice of positive ions held together by the sea of electrons, leading to stable structures.

Ionic and Metallic Comps.

3-page Review +

39 Question Study Guide

Chemistry Metallic Bonds Study Guide Answers

Marco Cascella



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